

# Simple Stoichiometry Lab

Name \_\_\_\_\_

Period \_\_\_\_\_

Date \_\_\_\_\_

Objective: To calculate the theoretical yield in a reaction, perform the reaction, and calculate the percent yield.

Background: In class, we've been talking about stoichiometry, the prediction of reaction masses using chemical equations. In this lab, you will be mixing acetic acid with sodium hydroxide to form sodium acetate and water. The equation for this reaction is:



You will start with a known quantity of acetic acid and sodium hydroxide and will predict the mass of sodium acetate formed. You will then measure the amount of sodium acetate actually produced by boiling the reaction mixture to dryness. By calculating your percent yield, you can determine your accuracy!

Pre-lab Questions:

1. Is boiling the water to isolate the sodium acetate the best way to measure the amount of sodium acetate formed? Explain why or why not.

Data:

Mass of weighed NaOH in 1. reaction	
Mass of weighed acetic	
Mass of sodium acetate produced	
Mass of NaOH in 1. reaction, after reaction is complete	

Sample Questions:

For answers to these questions see YOCUP at [www.yocup.com](http://www.yocup.com) or ask your instructor about where to go to get the answers.

1. Under the equation below, write the measured mass that you used for each reactant as given during the lab.



2. Using your measured mass of the acetic acid, calculate the theoretical yield of sodium acetate. (Show your work, factor-label method!)

3. Calculate the actual yield of sodium acetate using the values in the data section.

4. Calculate the percent yield of sodium acetate. Show your work!

5. According to the Law of Conservation of Mass, we cannot create atoms. Therefore, it is impossible to achieve more than 100% yield. What possible experimental errors could have occurred to cause some lab groups to have more than 100% yield? Think! Discuss error is NOT an acceptable answer!

6. It is highly unlikely that you would ever achieve 100% yield. Most often, lab groups fall short of 100%. What possible experimental errors could have occurred to decrease yield? Think! Not because error!

# Stoichiometry Lab Graph

**SA Dillow**



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