

Multiple-Choice Questions

Answer the following questions in 25 minutes. You may not use a calculator. You may use the periodic table and the equation sheet at the back of this book.

1. How many milliliters of 0.100 *M* H_2SO_4 are required to neutralize 50.0 mL of 0.200 *M* KOH ?
(A) 25.0 mL
(B) 30.0 mL
(C) 20.0 mL
(D) 50.0 mL
2. A sample of oxalic acid, $\text{H}_2\text{C}_2\text{O}_4$, is titrated with standard sodium hydroxide, NaOH , solution. A total of 45.20 mL of 0.1200 *M* NaOH is required to completely neutralize 20.00 mL of the acid. What is the concentration of the acid?
(A) 0.2712 *M*
(B) 0.1200 *M*
(C) 0.1356 *M*
(D) 0.2400 *M*
3. A solution is prepared by mixing 50.0 mL of 0.20 *M* arsenic acid, H_3AsO_4 , and 50.0 mL of 0.20 *M* sodium hydroxide, NaOH . Which anion is present in the highest concentration?
(A) HAsO_4^{2-}
(B) OH^-
(C) H_2AsO_4^-
(D) Na^+
4. $14 \text{H}^+(\text{aq}) + 6 \text{Fe}^{2+}(\text{aq}) + \text{Cr}_2\text{O}_7^{2-}(\text{aq}) \rightarrow 2 \text{Cr}^{3+}(\text{aq}) + 6 \text{Fe}^{3+}(\text{aq}) + 7 \text{H}_2\text{O}(\text{l})$
This reaction is used in the titration of an iron solution. What is the concentration of the iron solution if it takes 45.20 mL of 0.1000 *M* $\text{Cr}_2\text{O}_7^{2-}$ solution to titrate 50.00 mL of an acidified iron solution?
(A) 0.5424 *M*
(B) 0.1000 *M*
(C) 1.085 *M*
(D) 0.4520 *M*
5. Which of the following best represents the balanced net ionic equation for the reaction of silver carbonate with concentrated hydrochloric acid? Silver ions have a 1+ charge. In this reaction, all silver compounds are insoluble.
(A) $2 \text{AgCO}_3 + 4 \text{H}^+ + \text{Cl}^- \rightarrow \text{Ag}_2\text{Cl} + 2 \text{CO}_2 + 2 \text{H}_2\text{O}$
(B) $\text{Ag}_2\text{CO}_3 + 2 \text{H}^+ + 2 \text{Cl}^- \rightarrow 2 \text{AgCl} + \text{CO}_2 + \text{H}_2\text{O}$
(C) $\text{AgCO}_3 + 2 \text{H}^+ \rightarrow \text{Ag}^+ + \text{CO}_2 + \text{H}_2\text{O}$
(D) $\text{Ag}_2\text{CO}_3 + 4 \text{Cl}^- \rightarrow 2 \text{AgCl}_2 + \text{CO}_3^{2-}$

Stoichiometry Multiple Choice Questions

Singhal Atul



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Problems and Problem Solving in Chemistry Education Georgios Tsaparlis, 2021-05-17 Problem solving is central to the teaching

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