

## **Stoichiometry Practice Worksheet**

*Solve the following stoichiometry grams-grams problems:*

- 1) Using the following equation:  $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

- 2) Using the following equation:  $\text{Pb}(\text{SO}_4)_2 + 4 \text{ LiNO}_3 \rightarrow \text{Pb}(\text{NO}_3)_4 + 2 \text{ Li}_2\text{SO}_4$

How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

## **Limiting Reagent Worksheet**

*For the following reactions, find the following:*

- a) Which of the reagents is the limiting reagent?  
b) What is the maximum amount of each product that can be formed?  
c) How much of the other reagent is left over after the reaction is complete?

- 3) Consider the following reaction:



Answer the questions above (a,b,c), assuming we started with 30 grams of ammonium nitrate and 50 grams of sodium phosphate.

- 4) Consider the following reaction:



Answer the questions above (a,b,c) assuming we start with 100 grams of calcium carbonate and 45 grams of iron (III) phosphate.

5) When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

- a) Write the balanced equation for the reaction given above:  
b) If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?  
c) What is the limiting reagent for the reaction in #2? \_\_\_\_\_  
d) How much of the nonlimiting reagent is left over in this reaction?



# Stoichiometry Practice Chemfiesta Answers

**Sunil Tanna**





## **Stoichiometry Practice Chemfiesta Answers:**

Bite-Sized Chemistry Calculations John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension. Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide. This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield. The series is packed with many salient features that are meant to facilitate both teaching and learning. Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices. With this book you will be well prepared for your exams and boost your performance.

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Answers to Practice Problems

**Appendices**

**Sassy Stoichiometry Problems** Julie C. Gilbert, 2021-03-14 Need more Stoichiometry practice. Stoichiometry has been striking fear into the hearts of chemistry students for ages. The best way to conquer something is to practice it. Inside you'll find Brief descriptions of each type of ideal stoichiometry and limiting reactant stoichiometry 4 ideal stoichiometry worksheets broken down by type with keys and explanations 4 ideal stoichiometry self quizzes with their answer keys 2 limiting reactant stoichiometry worksheets with keys and explanations 2 limiting reactant stoichiometry self quizzes with answer keys 2 mixed stoichiometry self tests with answer keys. This is a companion workbook for the 5 Steps to Surviving Chemistry book. However you do not need to have read that book to find this workbook useful.

**Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry. As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all. The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities. For the serious student this means more often than not the frustrations of questions unanswered. The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox



reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction      **Stoichiometry** Sydney Young,1908      Chemistry Skill Builder Steven D. Gammon,1997 Chemistry SKILLBuilder gives students extra practice and feedback on three key topics nomenclature stoichiometry and balancing equations      **The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises      *A Stoichiometry Unit* David Callaghan,2004  
**Chemistry Success in 20 Minutes a Day** Michael B. McGinnis,2005 Offers a diagnostic test and twenty lessons covering vital chemistry skills      You Can Do Chemistry Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018



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