

Name: Key

Class: \_\_\_\_\_

Date: \_\_\_\_\_

Stoichiometry Problem Set A

**Directions:** Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

**Part 1: Straight Stoichiometry Problems**

**Target Goal:** To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

- 1.) For the reaction  $2\text{HNO}_3 + 1\text{Mg(OH)}_2 \rightarrow 1\text{Mg(NO}_3)_2 + 2\text{H}_2\text{O}$ , how many grams of magnesium <sup>B</sup> nitrate are produced from 8.00 mol of nitric acid ( $\text{HNO}_3$ )? A

$$\text{MM} = 148.3\text{g}$$

$$\frac{8.00\text{mol HNO}_3}{1} \times \frac{1\text{mol Mg(NO}_3)_2}{2\text{mol HNO}_3} \times \frac{148.3\text{g Mg(NO}_3)_2}{1\text{mol Mg(NO}_3)_2} = 592.8\text{g Mg(NO}_3)_2$$

- 2.) For the reaction  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ , how many moles of potassium <sup>B</sup> chlorate are required to produce 160. g of oxygen? A

$$\text{MM} = 32.0\text{g}$$

$$\frac{160.\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32.0\text{g O}_2} \times \frac{2\text{mol KClO}_3}{3\text{mol O}_2} = 3.33\text{ mole KClO}_3$$

- 3.) For the reaction  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ , how many grams of water are produced from 5.22 liters of oxygen gas? <sup>B</sup> A

$$\text{MM} = 18.0\text{g}$$

$$\frac{5.22\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{2\text{mol H}_2\text{O}}{1\text{mol O}_2} \times \frac{18.0\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 8.39\text{g H}_2\text{O}$$

- 4.) In the reaction  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how many grams of hydrogen <sup>B</sup> are produced if 1.33 grams of sodium react completely? A

$$\text{MM} = 23.0\text{g}$$

$$\text{MM} = 2.0\text{g}$$

$$\frac{1.33\text{g Na}}{1} \times \frac{1\text{mol Na}}{23.0\text{g Na}} \times \frac{1\text{mole H}_2}{2\text{mole Na}} \times \frac{2.0\text{g H}_2}{1\text{mole H}_2} = 0.0578\text{g H}_2$$

- 5.) In the reaction  $1\text{CO}_2 + 1\text{H}_2\text{O} \rightarrow 1\text{H}_2\text{CO}_3$ , how many moles of carbonic <sup>B</sup> acid can be produced from 97.2 g of  $\text{CO}_2$ ? A

$$\text{MM} = 44.0\text{g}$$

$$\frac{97.2\text{g CO}_2}{1} \times \frac{1\text{mole CO}_2}{44.0\text{g CO}_2} \times \frac{1\text{mole H}_2\text{CO}_3}{1\text{mole CO}_2} = 2.21\text{ moles H}_2\text{CO}_3$$

**Did you achieve our target goal?**

**Target Goal:** To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

1

Feeling VERY confused

2

Not sure that I get this

3

I think I am catching on

4

Doing great! Bring it on!

# Stoichiometry Problem Set Answers

**CO Houle**



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## **Table of Contents Stoichiometry Problem Set Answers**

1. Understanding the eBook Stoichiometry Problem Set Answers
  - The Rise of Digital Reading Stoichiometry Problem Set Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Problem Set Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Stoichiometry Problem Set Answers
  - User-Friendly Interface

4. Exploring eBook Recommendations from Stoichiometry Problem Set Answers
  - Personalized Recommendations
  - Stoichiometry Problem Set Answers User Reviews and Ratings
  - Stoichiometry Problem Set Answers and Bestseller Lists
5. Accessing Stoichiometry Problem Set Answers Free and Paid eBooks
  - Stoichiometry Problem Set Answers Public Domain eBooks
  - Stoichiometry Problem Set Answers eBook Subscription Services
  - Stoichiometry Problem Set Answers Budget-Friendly Options
6. Navigating Stoichiometry Problem Set Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Stoichiometry Problem Set Answers Compatibility with Devices
  - Stoichiometry Problem Set Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Stoichiometry Problem Set Answers
  - Highlighting and Note-Taking Stoichiometry Problem Set Answers
  - Interactive Elements Stoichiometry Problem Set Answers
8. Staying Engaged with Stoichiometry Problem Set Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Stoichiometry Problem Set Answers
9. Balancing eBooks and Physical Books Stoichiometry Problem Set Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Stoichiometry Problem Set Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Problem Set Answers
  - Setting Reading Goals Stoichiometry Problem Set Answers
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Stoichiometry Problem Set Answers
  - Fact-Checking eBook Content of Stoichiometry Problem Set Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web 2 4 2 instruction direction 42 2 4 3 valid architect s instruction 42 2 5 provisions empowering architect s instruction 44 2 5 1 pam 2006 44 2 5 2 other malaysian forms 48 2 5 3 international forms 40 2 5 4 compliance with architect s instruction 52 2 6 disputes of architect s instruction 52 2 7 summary 49

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