



SOLVING STOICHIOMETRY PROBLEMS

- a) mole A \rightarrow mole B
- b) mole A \rightarrow gram B
- c) gram A \rightarrow mole B
- d) gram A \rightarrow gram B

1. Find the molar mass of each compound present in the equation
2. Balance chemical equation
3. Find the given and identify the unknown
4. Refer to the diagram for sequence of formula

Ex:



MM (g/mol): (169.88) (208.23) (143.32) (261.35)

1. The no. of moles of BaCl_2 that could be produced from 4.25 moles of AgCl . (Mole to mole)

Mole A / known mole: 4.25 moles of AgCl
Mole B / unknown mole: x moles of BaCl_2

Mole A $\xrightarrow{\text{SE}}$ mole B

$$4.25 \text{ (3SF) moles AgCl} \times \frac{1 \text{ mole BaCl}_2}{2 \text{ moles AgCl}} = 2.13 \text{ (3SF) moles BaCl}_2$$

Coefficient in the balanced equation

2. The no. of grams of AgNO_3 that could be produced from 0.28 moles of $\text{Ba}(\text{NO}_3)_2$ (mole to grams)

Mole A / known : 0.28 moles of $\text{Ba}(\text{NO}_3)_2$
Gram B / unknown : x grams of AgNO_3

Mole A $\xrightarrow{\text{SE}}$ mole B $\xrightarrow{\text{MM}}$ gram B

$$0.28 \text{ (2SF) moles Ba(NO}_3)_2 \times \frac{2 \text{ mole AgNO}_3}{1 \text{ mole Ba(NO}_3)_2} \times \frac{169.88 \text{ g AgNO}_3}{1 \text{ mole AgNO}_3} = 95 \text{ (2SF) g AgNO}_3$$

MM of compound from equation

3. The no. of moles of AgCl that would be produced if 13.3 g of AgNO_3 are produced in the reaction. (gram to moles)

Gram A / known : 13.3 grams of AgNO_3
Mole B / unknown : x moles of AgCl

gram A $\xrightarrow{\text{MM}}$ mole A $\xrightarrow{\text{SE}}$ mole B

$$13.3 \text{ g (3SF) g AgCl} \times \frac{1 \text{ mole AgNO}_3}{169.88 \text{ g AgNO}_3} \times \frac{1 \text{ mole AgCl}}{2 \text{ mole AgNO}_3} = 0.040 \text{ (3 sf) g AgCl}$$

Stoichiometry Problem Solving Packet Holt Chemistry

Michell Joseph Sienko



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 - Chapter 2: Essential Elements of Stoichiometry Problem Solving Packet Holt Chemistry
 - Chapter 3: Stoichiometry Problem Solving Packet Holt Chemistry in Everyday Life
 - Chapter 4: Stoichiometry Problem Solving Packet Holt Chemistry in Specific Contexts
 - Chapter 5: Conclusion
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