

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar ratio to convert between moles of different substances.

Skill 1 - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill 2 - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[\frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[\frac{2.8\text{g}}{4.0\text{g}} \right] \times 100 = 70\%$$

Skill 3 - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol CaCO}_3 \times \frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} = 1.55\text{mol CaCl}_2$$

$$250\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} = 6.857\text{mol HCl} \quad \text{E.R.} \\ 6.857\text{mol HCl} \times \frac{1\text{mol CaCl}_2}{2\text{mol HCl}} = 3.43\text{mol CaCl}_2$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[\frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[\frac{142.0\text{g}}{172.02\text{g}} \right] \times 100 = 82.55\%$$

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A Stoichiometry Unit David Callaghan, 2004 **Stoichiometry Unit Project** Luann Marie Decker, 1998 **MCAT General Chemistry Review 2023-2024** Kaplan Test Prep, 2022-08-02 Kaplan's MCAT General Chemistry Review 2023 2024 offers an expert study plan detailed subject review and hundreds of online and in book practice questions all authored by the experts behind the MCAT prep course that has helped more people get into medical school than all other major courses combined Prepping for the MCAT is a true challenge Kaplan can be your partner along the way offering guidance on where to focus your efforts and how to organize your review This book has been updated to match the AAMC's guidelines precisely no more worrying about whether your MCAT review is comprehensive The Most Practice More than 350 questions in the book and access to even more online more practice than any other MCAT general chemistry book on the market The Best Practice Comprehensive general chemistry subject review is written by top rated award winning Kaplan instructors Full color 3 D illustrations from Scientific American charts graphs and diagrams help turn even the most complex science into easy to visualize concepts All material is vetted by editors with advanced science degrees and by a medical doctor Online resources including a full length practice test help you practice in the same computer based format you'll see on Test Day Expert Guidance High yield badges throughout the book identify the topics most frequently tested by the AAMC We know the test The Kaplan MCAT team has spent years studying every MCAT related document available Kaplan's expert psychometricians ensure our practice questions and study materials are true to the test **MCAT General Chemistry Review 2020-2021** Kaplan Test Prep, 2019-07-02 Kaplan's MCAT General Chemistry Review 2020 2021 is updated to reflect the latest most accurate and most testable materials on the MCAT A new layout makes our book even more streamlined and intuitive for easier review You'll get efficient strategies detailed subject review and hundreds of practice questions all authored by the experts behind the MCAT prep course that has helped more people get into medical school than all other major courses combined Efficient Strategies and In Depth Review High Yield badges indicate the most testable content based on AAMC materials Concept summaries that boil down the need to know information in each chapter including any necessary equations to memorize Chapter Profiles indicate the degree to which each chapter is tested and the testmaker content categories to which it aligns Charts graphs diagrams and full color 3 D illustrations from Scientific American help turn even the most complex science into easy to visualize concepts Realistic Practice One year online access to instructional videos practice questions and quizzes Hundreds of practice questions show you how to apply concepts and equations 15 multiple choice Test Your Knowledge questions at the end of each chapter Learning objectives and concept checks ensure you're focusing on the most important information in each chapter Expert Guidance Sidebars illustrate connections between concepts and include references to more information real world tie ins mnemonics and MCAT specific tips Comprehensive subject review written by top rated award winning Kaplan instructors who guide you on where to focus your efforts and how to organize your review

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