

Stoichiometry

Lecture Examples Answer Key

Ex. 1 Balance the following chemical equations:



Ex. 2 Write a balanced chemical reaction for the combustion of octane, $\text{C}_8\text{H}_{18(l)}$, in air.



Ex. 3 Write the balanced chemical equation for the combination of the metallic element calcium with the nonmetallic element oxygen, O_2 .



Ex. 4 Calculate the molecular weight of

(a) N_2O_5 **MW = 108.0104 g mol⁻¹**

(b) $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ **MW = 158.167 g mol⁻¹**

(c) $[\text{K}_2(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 3\text{H}_2\text{O}]$ **MW = 902.1758 g mol⁻¹**

Ex. 5 Calculate the percentage by mass of oxygen in 4a and 4b above.

(a) $[(5 \times 15.9994 \text{ g mol}^{-1}) / 108.0104 \text{ g mol}^{-1}] \times 100\% = \mathbf{74.0642\%}$

(b) $[(4 \times 15.9994 \text{ g mol}^{-1}) / 158.167 \text{ g mol}^{-1}] \times 100\% = \mathbf{40.4620\%}$

Ex. 6 (a) Calculate the number of O atoms in 0.470 mol of $\text{C}_6\text{H}_{12}\text{O}_6$.

$$\begin{aligned} \# \text{ O atoms} &= (0.470 \text{ mol C}_6\text{H}_{12}\text{O}_6)(6.022 \times 10^{23} \text{ molecules/1 mol})(6 \text{ O atoms / 1 molecule}) \\ &= \mathbf{1.70 \times 10^{24} \text{ O atoms}} \end{aligned}$$

(b) Calculate the total number of ions in 38.1 g of CaF_2 .

$$\begin{aligned} \# \text{ ions} &= (38.1 \text{ g CaF}_2)(1 \text{ mol / 78.075 g})(6.022 \times 10^{23} \text{ molecules/1 mol})(3 \text{ ions / 1 molecule}) \\ &= \mathbf{8.82 \times 10^{23} \text{ ions}} \end{aligned}$$

Ex. 7 (a) What is the mass in grams of 1 mol of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?

$$\text{mass} = (1 \text{ mol})(180.1572 \text{ g/1 mol}) = \mathbf{180.1572 \text{ g}}$$

(b) What is the mass in grams of 3.52 mol of chromium(III) sulfate decahydrate?

Stoichiometry Year 12 Questions

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A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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Chemistry Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced eText with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order

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