

Stoichiometry Study Guide Name: Angie Julieth Rodriguez Angel

Mole Concept questions

- 1) What SI unit is used to measure the number of representative particles in a substance? **The mole.**
- 2) What is Avogadro's number? **6.023×10^{23}**
- 3) What is the first step in all stoichiometry problems? **Grams and moles.**
- 4) What is a limiting reagent? **Is the reactant that gets consumed first in a chemical reaction and therefore limits how much product can be formed.**
- 5) What is an excess reagent? **Is the reactant that could keep reacting if the other had not been consumed.**
- 6) Define: theoretical yield, actual yield, % yield.
 - **theoretical yield:** what the balanced equation tells should be made (your calculations).
 - **Actual yield:** What you actually get in the lab when the chemicals are mixed.
 - **Yield:** is the amount of product made in a chemical reaction.
- 7) Which will always be larger (theoretical yield or actual yield)? **Is almost always smaller than the theoretical yield.**
- 8) The calculation of quantities in chemical equations is called **stoichiometry**
- 9) What determines the molar ratio between products and reactants? **the coefficients in the balanced chemical equation.**
- 10) What do we do that shows that mass, moles and atoms are conserved? **all atoms present within the reactants are conserved, and must be present within the products when the reaction is over.**

Molar mass and percent composition:

- 11) What is the molar mass of sucrose ($C_{12}H_{22}O_{11}$)? **342g/mol**
- 12) What is the molar mass of each of the following compounds? Hint write the formula by either criss-crossing charges (ionic) or using the prefix (covalent) given.
 - a. phosphorus pentachloride **covalent**
 - b. uranium hexafluoride (UF_6) **ionic**
 - c. Calcium phosphate **ionic**
- 13) What is the % composition of potassium in KBr? **32.66%**
- 14) Type your name here: **Angie Julieth Rodriguez Angel**

Looking at mole ratios:

- 15) How many total moles are in the **reactants** in the equation $2Na + Cl_2 \rightarrow 2NaCl$? **Sodium metal, Na, chlorine gas.**
- 16) How many total moles of **products** are made in the equation $C_2H_2 + 8O_2 \rightarrow 5CO_2 + 6H_2O$ **no se.**
- 17) Using the equation above, what is the mole ratio of oxygen to water? **1:2**
- 18) Using the equation above, what is the mole ratio of C_2H_2 to carbon dioxide? **5 moles of CO2.**

Balance the following $Fe(III) + O_2 \rightarrow Fe_2O_3$

- 19) Using the equation above, What is the molar ratio between Fe and O_2 **5.5 mol g**
- 20) Using the equation above, what is the molar ratio between Fe_2O_3 and O_2 **4:2**

Stoichiometry Chapter 1study Guide

Arshad Iqbal



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