

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) =14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) +4(2) =14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) =6 O atoms

Product side

4 Fe atoms

2(3) =6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃:4 mol Fe

Stoichiometry Introduction Packet Answers

NARAYAN CHANGDER



Stoichiometry Introduction Packet Answers:

Russian Chemical Reviews ,1986 **Subject Guide to Forthcoming Books** ,1983 Presents by subject the same titles that are listed by author and title in Forthcoming books **An Introduction to Stoichiometry** ,2016 What are the quantitative relationships between the substances in a chemical reaction The study of stoichiometry shows you how to apply your ability to balance equations to solve problems involving chemical reactions **Stoichiometry ... Together with an Introduction to the Study of Physical Chemistry, by Sir William Ramsay** Sydney YOUNG,1908 **Calculating in Chemistry** Cosmas C Mbuji,2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry Introduction to Process Calculations Stoichiometry KA. Gavhane,2012

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Guided Inquiry for General Chemistry (First Edition) J. Hugh Broome,Mary Mackey,2020-12-31 Guided Inquiry for General Chemistry provides students with an interactive introduction to key concepts in chemistry This workbook covers all of the topics and ideas presented within a first year chemistry course for science majors Short chapters guide students to understanding through simple questions followed by more advanced practice exercises designed to be completed in a group setting with instructor assistance Each chapter introduces readers to fundamental chemistry concepts challenges them to think and reflect on those concepts and examines essential applications of those concepts Topics in the book include atomic structure bonding Lewis dot structures nomenclature chemical reaction types stoichiometry states of matter kinetics equilibrium energetics electrochemistry and nuclear chemistry Each chapter features explicitly stated learning outcomes a list of prerequisite chapters that will assist readers in their understanding of the current chapter background information with guiding questions and application questions to facilitate learning and retention Comprehensive and approachable in nature Guided Inquiry for General Chemistry is designed for first year chemistry courses at the university level but is also well suited for introductory and high school chemistry courses

Stoichiometry B I. Bhatt,2010

Friendly Chemistry Student Workbook Joey A

Hajda, Lisa B Hajda, 2021-02 Friendly Chemistry is a truly unique approach to teaching introductory chemistry Used by home schoolers and charter public and private school students world wide for over ten years Friendly Chemistry presents what is often considered an intimidating subject as a genuinely fun enjoyable experience Whether you re a high school aged student needing a lab science course or a non traditional student looking for a refresher course to help you prepare for an upcoming entrance exam Friendly Chemistry can help you accomplish your goal in a painless way If you do have aspirations of a future in a science field Friendly Chemistry can give you the solid foundation you need to succeed in subsequent courses Friendly Chemistry was written using simple language and a host of analogies to make learning and teaching chemistry easy The chemistry concepts presented in Friendly Chemistry are NOT watered down The concepts are just explained in ways that are readily understood by most learners Coupled with these explanations is a host of teaching aids labs and games which makes the learning concrete and multi sensory Students find the course fun and painless Parents often comment I wish I had had this when I was taking chemistry Now it all makes so much sense Friendly Chemistry covers the same topics taught in traditional high school chemistry courses The course begins with an introduction to atomic theory followed by discussion of why the elements are arranged the way they are in the periodic table Quantum mechanics comes next using the acclaimed Doo wop Board as a teaching aid Next comes a discussion of how atoms become charged ionization followed by an explanation of how charged atoms make compounds The mole is introduced next followed by a discussion of chemical reactions Stoichiometry predicting amounts of product produced from a reaction is treated next followed by a discussion of solutions molarity The course is wrapped up with a discussion of the ideal gas laws Please note that this is the STUDENT WORKBOOK This volume contains worksheets and lab report pages which accompany the student edition There is no text or other explanatory material in this workbook The student edition must be purchased separately More information regarding Friendly Chemistry including answers to many frequently asked questions may be found at www.friendlychemistry.com

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Introductory College Chemistry Students' Understanding of Stoichiometry Adam J. Wolfer, 2000 Previous research has shown a gap between chemistry students conceptual and computational understandings of chemistry topics such as gas laws equilibrium and reactions This qualitative study examined the conceptual and computational understandings of stoichiometry of college students enrolled in a large lecture Introductory Chemistry course Factors that might influence students understandings were examined to determine their influence Possible influential factors examined included students prior coursework and their current chemistry instruction Instruction on stoichiometry was examined through classroom observations an instructor interview and review of the course resources Course exams and out of class assignments were also examined for their influence on students Student volunteers n 6 were interviewed to gauge their understanding of stoichiometry Students understanding was assessed through tasks that included a card sort solving conceptual and computational problems drawing representations of reactions and answering questions concerning their philosophy of learning chemistry Results indicated that students had an acceptable understanding of the particulate nature of matter but did not apply this knowledge to problem solving The students were most comfortable solving computational problems where they could apply algorithms learned from their instructor The students also applied algorithms in answering conceptual problems There appeared to be a connection between the students conceptual structures of stoichiometry and their ability to solve computational problems The lack of conceptual questions in assessment appeared to be a major contributing factor in the students lack of conceptual understanding because the students discounted the importance of learning aspects of stoichiometry that were not included on exams Other contributing factors included the computational focus of instruction on limiting reactant problems textbook presentation and student exercises

Formulation and Stoichiometry Emil J Margolis,

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