

9.2 – Stoichiometry – Limiting Reagent



20 grams of copper (II) chloride react with 25 grams of sodium nitrate. Identify the limiting reagent.

$$20\text{g } \cancel{\text{CuCl}_2} \times \frac{1 \text{ mol } \cancel{\text{CuCl}_2}}{137.452\text{g } \cancel{\text{CuCl}_2}} = 0.149 \text{ mol } \cancel{\text{CuCl}_2} \times \frac{2 \text{ mol } \text{NaNO}_3}{1 \text{ mol } \cancel{\text{CuCl}_2}} = 0.298 \text{ mol } \text{NaNO}_3$$

HAVE Need

$$25\text{g } \cancel{\text{NaNO}_3} \times \frac{1 \text{ mol } \cancel{\text{NaNO}_3}}{84.994\text{g } \cancel{\text{NaNO}_3}} = 0.294 \text{ mol } \cancel{\text{NaNO}_3} \times \frac{1 \text{ mol } \text{CuCl}_2}{2 \text{ mol } \cancel{\text{NaNO}_3}} = 0.147 \text{ mol } \text{CuCl}_2$$

You need 0.298 mol NaNO_3 , but you have 0.294 mol NaNO_3

Stoichiometry Limiting Reagent Chemistry If8766

Answers

Sydney Young



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Introductory College Chemistry Students' Understanding of Stoichiometry Adam J. Wolfer,2000 Previous research has shown a gap between chemistry students' conceptual and computational understandings of chemistry topics such as gas laws equilibrium and reactions This qualitative study examined the conceptual and computational understandings of stoichiometry of college students enrolled in a large lecture Introductory Chemistry course Factors that might influence students' understandings were examined to determine their influence Possible influential factors examined included students' prior coursework and

their current chemistry instruction Instruction on stoichiometry was examined through classroom observations an instructor interview and review of the course resources Course exams and out of class assignments were also examined for their influence on students Student volunteers n 6 were interviewed to gauge their understanding of stoichiometry Students understanding was assessed through tasks that included a card sort solving conceptual and computational problems drawing representations of reactions and answering questions concerning their philosophy of learning chemistry Results indicated that students had an acceptable understanding of the particulate nature of matter but did not apply this knowledge to problem solving The students were most comfortable solving computational problems where they could apply algorithms learned from their instructor The students also applied algorithms in answering conceptual problems There appeared to be a connection between the students conceptual structures of stoichiometry and their ability to solve computational problems The lack of conceptual questions in assessment appeared to be a major contributing factor in the students lack of conceptual understanding because the students discounted the importance of learning aspects of stoichiometry that were not included on exams Other contributing factors included the computational focus of instruction on limiting reactant problems textbook presentation and student exercises

Stoichiometry Sydney Young, 1918 This introduction was originally prefixed to Dr Findlay s Phase rule which was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii Introduction issued also separately in 1904

Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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Formulation and Stoichiometry Emil J Margolis, **Improving Student Comprehension of Stoichiometric Concepts** Connie Lynn Bannick Kemner, 2007 **Chemistry Problems** Michell J. Sienko, 1967 Stoichiometry B I. Bhatt, 2010

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