

Stoichiometry Limiting Reagent

Answers

- 1 Consider the following reaction:



- (a) How many grams of NH_3 can be produced from the reaction of 28 g of N_2 and 25 g of H_2 ?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 34 \text{ g NH}_3$$

$$25 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2 \text{ g H}_2} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 142 \text{ g NH}_3$$

34 g of NH_3 can be produced.

- (b) How much of the excess reagent is left over?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} \times \frac{2 \text{ g H}_2}{1 \text{ mol H}_2} = 6 \text{ g H}_2 \text{ used}$$

$$25 \text{ g H}_2 - 6 \text{ g H}_2 = 19 \text{ g H}_2 \text{ remains}$$

- 2 Consider the following reaction:



What volume of H_2 at STP is produced from the reaction of 50 g of Mg and 75 g of HCl ?

$$50 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 46 \text{ L H}_2$$

$$75 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HCl}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 23 \text{ L H}_2$$

23 L H_2 is produced.

- 3 Zinc and sulfur react to form zinc sulfide according to the following equation:



Suppose 25 g of zinc and 30 g of sulfur react. Which reagent is limiting?

$$25 \text{ g Zn} \times \frac{1 \text{ mol Zn}}{65.4 \text{ g Zn}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol Zn}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 37.3 \text{ g ZnS}$$

$$30 \text{ g S} \times \frac{1 \text{ mol S}}{32 \text{ g S}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol S}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 91.2 \text{ g ZnS}$$

Zn is the limiting reagent.

Stoichiometry Limiting Reagent Worksheet Answers

Chemistry If8766

Cosmas C Mbuzi



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Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may

also find this book a good resource for their lessons in stoichiometry *You Can Do Chemistry* Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 **Sassy Stoichiometry Problems** Julie C. Gilbert,2021-03-14 Need more Stoichiometry practice Stoichiometry has been striking fear into the hearts of chemistry students for ages The best way to conquer something is to practice it Inside you ll find Brief descriptions of each type of ideal stoichiometry and limiting reactant stoichiometry 4 ideal stoichiometry worksheets broken down by type with keys and explanations 4 ideal stoichiometry self quizzes with their answer keys 2 limiting reactant stoichiometry worksheets with keys and explanations 2 limiting reactant stoichiometry self quizzes with answer keys 2 mixed stoichiometry self tests with answer keys This is a companion workbook for the 5 Steps to Surviving Chemistry book However you do not need to have read that book to find this workbook useful **Stoichiometry Unit Project** Luann Marie Decker,1998 **Stoichiometry** Sydney Young,1918 This introduction was originally prefixed to Dr Findlay s Phase rule which was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii Introduction issued also separately in 1904

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