

Stoichiometry: Balanced Problems (K&N)

- 11) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 Calculate volume of H_2O gas at STP is produced if 10 grams of H_2 gas reacted with an excess of O_2

$$\frac{2 \text{ moles } \text{H}_2}{22.4 \text{ L}} \times \frac{10 \text{ g } \text{H}_2}{2.0 \text{ g } \text{H}_2} \times \frac{2 \text{ moles } \text{H}_2\text{O}}{2 \text{ moles } \text{H}_2} \times \frac{22.4 \text{ L}}{1 \text{ mole } \text{H}_2\text{O}} = 112 \text{ L } \text{H}_2\text{O}$$

- 12) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 22.4 L of H_2O gas decomposed to form volume of H_2 is produced at STP

$$\frac{2 \text{ moles } \text{H}_2\text{O}}{22.4 \text{ L}} \times \frac{22.4 \text{ L } \text{H}_2\text{O}}{2 \text{ moles } \text{H}_2\text{O}} \times \frac{2 \text{ moles } \text{H}_2}{2 \text{ moles } \text{H}_2\text{O}} \times \frac{22.4 \text{ L}}{1 \text{ mole } \text{H}_2} = 22.4 \text{ L } \text{H}_2$$

- 13) 2 moles oxygen gas at STP are produced in Problem 11

$$\frac{2 \text{ moles } \text{H}_2\text{O}}{22.4 \text{ L}} \times \frac{10 \text{ g } \text{H}_2}{2.0 \text{ g } \text{H}_2} \times \frac{2 \text{ moles } \text{H}_2}{2 \text{ moles } \text{H}_2\text{O}} \times \frac{22.4 \text{ L}}{1 \text{ mole } \text{H}_2} = 22.4 \text{ L } \text{H}_2$$

- 14) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ What is O_2 ?
 Calculate volume of oxygen gas at STP is produced when 2.0 g of H_2 reacts with excess O_2 .
 Equations below used?

$$\frac{2 \text{ moles } \text{H}_2}{22.4 \text{ L}} \times \frac{2.0 \text{ g } \text{H}_2}{2.0 \text{ g } \text{H}_2} \times \frac{1 \text{ mole } \text{O}_2}{2 \text{ moles } \text{H}_2} \times \frac{22.4 \text{ L}}{1 \text{ mole } \text{O}_2} = 1.12 \text{ L } \text{O}_2$$

- 15) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ What is H_2O ?
 2 moles oxygen molecules is reacted and produced is 2 moles water molecules are produced. How volume is produced?

$$\frac{2 \text{ moles } \text{H}_2\text{O}}{22.4 \text{ L}} \times \frac{2.0 \text{ g } \text{H}_2}{2.0 \text{ g } \text{H}_2} \times \frac{2 \text{ moles } \text{H}_2\text{O}}{2 \text{ moles } \text{H}_2} \times \frac{22.4 \text{ L}}{1 \text{ mole } \text{H}_2\text{O}} = 2.24 \text{ L } \text{H}_2\text{O}$$

- 16) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 22.4 L of water molecules are the compound. How many molecules of H_2 are produced?

$$\frac{2 \text{ moles } \text{H}_2}{22.4 \text{ L}} \times \frac{22.4 \text{ L } \text{H}_2\text{O}}{2 \text{ moles } \text{H}_2\text{O}} \times \frac{2 \text{ moles } \text{H}_2}{2 \text{ moles } \text{H}_2\text{O}} \times \frac{22.4 \text{ L}}{1 \text{ mole } \text{H}_2} = 22.4 \text{ L } \text{H}_2$$

Stoichiometry Mixed Problems Key

M Tight



Stoichiometry Mixed Problems Key:

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be 22 414 l 22 414 l we can substitute 101 325 kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k for temperature into the ideal gas equation and solve for r r

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introduction the ideal gas law is a commonly used formula in chemistry it relates pressure temperature and volume this
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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between
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declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95
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