

## Quest: Stoichiometry 101

D3. Demonstrate an understanding of the mole concept and its significance to the quantitative analysis of chemical reactions.

R	1	2	3	4
Does not meet expectations	demonstrates <b>limited</b> knowledge and understanding of content	demonstrates <b>some</b> knowledge and understanding of content	demonstrates <b>considerable</b> knowledge and understanding of content	demonstrates <b>thorough</b> knowledge and understanding of content
does not meet expectations	uses critical/ creative thinking processes, skills, and strategies with <b>limited</b> effectiveness	uses critical/creative thinking processes, skills, and strategies with <b>some</b> effectiveness	uses critical/ creative thinking processes, skills, and strategies with <b>considerable</b> effectiveness	uses critical/ creative thinking processes, skills, and strategies with a <b>high</b> degree of effectiveness

**Instructions:** *For each of the calculations below be sure to show all your work and to round final answers to the correct number of significant digits.*

1. Propane burns in excess oxygen according to the following equation:



What mass of water is produced when 25.00 g of propane,  $\text{C}_3\text{H}_8(g)$  undergoes combustion?

First, we need to find the moles of propane burned:

$$\begin{aligned} m(\text{C}_3\text{H}_8) &= 25.00 \text{ g} \\ M(\text{C}_3\text{H}_8) &= 44.1 \text{ g/mol} \end{aligned}$$

$$n(\text{C}_3\text{H}_8) = m(\text{C}_3\text{H}_8) / M(\text{C}_3\text{H}_8) = 25.00 \text{ g} / 44.1 \text{ g/mol} = 0.566 \text{ mol}$$

From the balanced equation, we see that 1 mole of propane reacts with 4 moles of water:



So the moles of water produced can be calculated as:

$$n(\text{H}_2\text{O}) = 4 \times n(\text{C}_3\text{H}_8) = 4 \times 0.566 \text{ mol} = 2.264 \text{ mol}$$

Finally, we can use the molar mass of water to convert from moles to grams:

$$M(\text{H}_2\text{O}) = 18.0 \text{ g/mol}$$

$$m(\text{H}_2\text{O}) = n(\text{H}_2\text{O}) \times M(\text{H}_2\text{O}) = 2.264 \text{ mol} \times 18.0 \text{ g/mol} = 40.75 \text{ g}$$

Therefore, when 25.00 g of propane is burned, 40.75 g of water is produced.

# Stoichiometry Quest Answers

**LP Steffe**



## **Stoichiometry Quest Answers:**

## Unveiling the Magic of Words: A Report on "**Stoichiometry Quest Answers**"

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