ANSWERS:

Stoichiometry Quiz SCH3U Name:

___T __C __A Assessment of Learning (Product) 4 6

2 C₁·H₂··· + 29 O₁··· (excess) → 20 OO₁ + 18 H₂O

 Answer the following questions based on the balanced chemical equation above.

a) If 23.5 g of C_{1.0}H_m is mised with 3.76 s 10th molecules of oxygen, determine the limiting reactant, (show all work) (6T, 1C)

ANS: oxygen is limiting reactant.

b) What mass of H₂O is expected in part 42 (1T, 1C)

ANS: 6.99 g water

c) Now many molecules CO_t of are expected in part at (3T, 1C).

ANS: 2.59 x 16⁽²⁾ molecules of CO₁

d) What mass of the excess reactant is left over after the reaction! (3T, 1C)

ANS: 17.5 g Callin.

2. During a lab a student reacts 16.4 g of nitrogen with excess owners.

a. Balance the chemical equation above, (2A)

b. Calculate the molecules of drastrogen pentagoide that form, if the reaction proceeds with a 58.2% yield (4A)

ANS: 2.30 x 10¹² molecules of N₂O₄

Stoichiometry Quiz For Chemistry With Answers

Arshad Iqbal

Stoichiometry Quiz For Chemistry With Answers:

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