

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar/mole ratio to convert between moles of different substances.

Skill 1 - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill 2 - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{2.8\text{g}/4.0\text{g}}{1} \right] \times 100 = 70\%$$

Skill 3 - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol} \\ 250\text{g HCl} \times \frac{1\text{mol HCl}}{36.45\text{g HCl}} = 6.857\text{mol HCl} \quad \text{2mol in rxn} = 3.43\text{mol}$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{142.0\text{g}/172.02\text{g}}{1} \right] \times 100 = 82.55\%$$

Stoichiometry Test B Answer Key

E Durkheim



Stoichiometry Test B Answer Key:

The book delves into Stoichiometry Test B Answer Key. Stoichiometry Test B Answer Key is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Stoichiometry Test B Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Stoichiometry Test B Answer Key
- Chapter 2: Essential Elements of Stoichiometry Test B Answer Key
- Chapter 3: Stoichiometry Test B Answer Key in Everyday Life
- Chapter 4: Stoichiometry Test B Answer Key in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Stoichiometry Test B Answer Key. This chapter will explore what Stoichiometry Test B Answer Key is, why Stoichiometry Test B Answer Key is vital, and how to effectively learn about Stoichiometry Test B Answer Key.
3. In chapter 2, the author will delve into the foundational concepts of Stoichiometry Test B Answer Key. This chapter will elucidate the essential principles that must be understood to grasp Stoichiometry Test B Answer Key in its entirety.
4. In chapter 3, this book will examine the practical applications of Stoichiometry Test B Answer Key in daily life. The third chapter will showcase real-world examples of how Stoichiometry Test B Answer Key can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Stoichiometry Test B Answer Key in specific contexts. The fourth chapter will explore how Stoichiometry Test B Answer Key is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Stoichiometry Test B Answer Key. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Stoichiometry Test B Answer Key.

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