

4 • Chemical Equations and Stoichiometry**STUDY QUESTIONS AND PROBLEMS**

1. Balance the following equations:

- a. $__\text{C}_4\text{H}_8(\text{g}) + __\text{O}_2(\text{g}) \rightarrow __\text{CO}_2(\text{g}) + __\text{H}_2\text{O}(\text{l})$
- b. $__\text{NH}_3(\text{g}) + __\text{O}_2(\text{g}) \rightarrow __\text{NO}_2(\text{g}) + __\text{H}_2\text{O}(\text{l})$
- c. $__\text{PCl}_3(\text{l}) + __\text{H}_2\text{O}(\text{l}) \rightarrow __\text{H}_3\text{PO}_3(\text{aq}) + __\text{HCl}(\text{aq})$
- d. $__\text{Ca}_3\text{P}_2(\text{s}) + __\text{H}_2\text{O}(\text{l}) \rightarrow __\text{Ca}(\text{OH})_2(\text{aq}) + __\text{PH}_3(\text{g})$
- e. $__\text{C}_4\text{H}_8(\text{OH})_2(\text{l}) + __\text{O}_2(\text{g}) \rightarrow __\text{CO}_2(\text{g}) + __\text{H}_2\text{O}(\text{l})$
- f. $__\text{NH}_3(\text{g}) + __\text{NO}(\text{g}) \rightarrow __\text{N}_2(\text{g}) + __\text{H}_2\text{O}(\text{l})$
- g. $__\text{KClO}_3(\text{s}) \rightarrow __\text{KCl}(\text{s}) + __\text{O}_2(\text{g})$
- h. $__\text{Ca}(\text{OH})_2(\text{s}) + __\text{H}_3\text{PO}_4(\text{aq}) \rightarrow __\text{Ca}_3(\text{PO}_4)_2(\text{s}) + __\text{H}_2\text{O}(\text{l})$
- i. $__\text{C}_3\text{H}_8(\text{g}) + __\text{O}_2(\text{g}) \rightarrow __\text{CO}_2(\text{g}) + __\text{H}_2\text{O}(\text{l})$
- j. $__\text{N}_2\text{O}(\text{g}) + __\text{O}_2(\text{g}) \rightarrow __\text{NO}_2(\text{g})$
- k. $__\text{Al}_4\text{C}_3(\text{s}) + __\text{H}_2\text{O}(\text{l}) \rightarrow __\text{Al}(\text{OH})_3(\text{aq}) + __\text{CH}_4(\text{g})$
- l. $__\text{CS}_2(\text{l}) + __\text{Cl}_2(\text{g}) \rightarrow __\text{CCl}_4(\text{l}) + __\text{S}_2\text{Cl}_2(\text{l})$
- m. $__\text{C}_2\text{H}_5\text{OH}(\text{l}) + __\text{PCl}_3(\text{l}) \rightarrow __\text{C}_2\text{H}_5\text{Cl}(\text{l}) + __\text{H}_3\text{PO}_3(\text{l})$
- n. $__\text{ZnS}(\text{s}) + __\text{O}_2(\text{g}) \rightarrow __\text{ZnO}(\text{s}) + __\text{SO}_2(\text{g})$
- o. $__\text{Ag}(\text{s}) + __\text{H}_2\text{S}(\text{g}) + __\text{O}_2(\text{g}) \rightarrow __\text{Ag}_2\text{S}(\text{s}) + __\text{H}_2\text{O}(\text{l})$

3. When asked to balance the equation:



the following suggestions were made:



Which answer is correct and what is wrong with the others?

4. Write balanced chemical equations for the following reactions:

- a. the decomposition of ammonium nitrate to nitrogen gas, oxygen gas, and water vapor.
- b. the reaction of sodium bicarbonate with sulfuric acid to produce sodium sulfate, water, and carbon dioxide.
- c. the treatment of phosphorus pentachloride with water to produce phosphoric acid and hydrogen chloride.

Stoichiometry The Arithmetic Of Equations Answer Key

Yan Bai



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