

# Stoichiometry Study Guide

Balance the following equations. Then use the balanced equation to answer questions #1-5.

1) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

2) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

Balance the following equation.

3) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

4) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## Stoichiometry Study Guide

Fill in the blanks below.

1 mole = \_\_\_\_\_ grams.

Balance the following equation. Then use the balanced equation to answer the questions #1-5.



1) Find the formula for the following compounds:

a) NaCl to  $\text{O}_2$

b) NaCl to  $\text{NaClO}_2$

c)  $\text{O}_2$  to  $\text{NaClO}_2$

2) If the above chemical reaction is done as an experiment using 4.4 moles of each reactant, calculate how much  $\text{NaClO}_2$  will be produced?

a) Given 4.4 moles of NaCl, how many moles of  $\text{NaClO}_2$  will be produced?

b) Given 4.4 moles of  $\text{O}_2$ , how many moles of  $\text{NaClO}_2$  will be produced?

c) Define limiting reactant & excess reactant.

d) Which reactant is the limiting reactant? Which reactant is the excess reactant?

3) If the above chemical reaction is done as an experiment using 10 moles of  $\text{O}_2$ , calculate the percent yield given that the actual yield of  $\text{NaClO}_2$  is 8.8 moles.

a) What is the theoretical yield in grams of  $\text{NaClO}_2$  given 10 moles of the  $\text{O}_2$ ?

b) Write the formula for Percent Yield.

c) What is the percent yield given that the actual yield of  $\text{NaClO}_2$  is 8.8 moles?

# Stoichiometry Unit Study Guide

**Mark Steyvers**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Stoichiometry Unit Study Guide**

1. Understanding the eBook Stoichiometry Unit Study Guide
  - The Rise of Digital Reading Stoichiometry Unit Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Unit Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Stoichiometry Unit Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Unit Study Guide
  - Personalized Recommendations
  - Stoichiometry Unit Study Guide User Reviews and Ratings
  - Stoichiometry Unit Study Guide and Bestseller Lists
5. Accessing Stoichiometry Unit Study Guide Free and Paid eBooks
  - Stoichiometry Unit Study Guide Public Domain eBooks
  - Stoichiometry Unit Study Guide eBook Subscription Services

- Stoichiometry Unit Study Guide Budget-Friendly Options
- 6. Navigating Stoichiometry Unit Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Stoichiometry Unit Study Guide Compatibility with Devices
  - Stoichiometry Unit Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Stoichiometry Unit Study Guide
  - Highlighting and Note-Taking Stoichiometry Unit Study Guide
  - Interactive Elements Stoichiometry Unit Study Guide
- 8. Staying Engaged with Stoichiometry Unit Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Stoichiometry Unit Study Guide
- 9. Balancing eBooks and Physical Books Stoichiometry Unit Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Stoichiometry Unit Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Unit Study Guide
  - Setting Reading Goals Stoichiometry Unit Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Unit Study Guide
  - Fact-Checking eBook Content of Stoichiometry Unit Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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