Chemistry 12



b) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of consumption of HCI in moles/s.

5.0×10-9 mal Mg × 2 mal HC1 = 1.0×10-8 mal HC1/s

Answer 1.0×10-8 mal HC1/s

c) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of production of H₂ in g/s.

5.0x10-9 mol Mg x 1 mol Hz x 20 g Hz = 1.0x10-8 g Hz/s

d) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of production of H₂ in L/s (@STP).

H2 in L/s (@STP).

5.0 x 10 - 9 mol Mg x 1 mol H2 x 22.4 L H2 = Answer 1.1 x 10 - 7 L H2/s

 e) If the rate of consumption of magnesium is 5.0 x 10-9 mol/s, find the mass of Mg consumed in 5.0 minutes.

5.0×10-9 mal Mg × 21.3 g Mg × 60 × 5.0 min = 3.6×10-5 g Mg

- When butane (C₄H₁₀) is burned in air (oxygen), the products carbon dioxide and water are formed.
 - a) Write a balanced formula equation for this reaction.

2 C4H10 +1302 -> 8CO2 + 10H20

b) If butane is consumed at an average rate of 0.116 grams/s, determine the rate of production of CO₂ in g/s.

0.116 g Cathio x 1 mol Cathio x 8 mol CO2 x 49.0 g CO2 = 0.352 g CO2/s

Answer 0.352 g CO2/s

Stoichiometry Using Copper Lab 31 Answer Key

Karl Heinz Weiss, Michael Schilsky

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Wilson Disease Karl Heinz Weiss, Michael Schilsky, 2019-05-22 Wilson Disease Pathogenesis Molecular Mechanisms Diagnosis Treatment and Monitoring translates both clinical and experimental findings into a comprehensive approach for anyone involved in research and patient care While the clinical variability of Wilson Disease poses a challenge from a diagnostic approach the book uses the translational impact of new research findings to relate to new treatment concepts Comprehensive chapters include common knowledge guideline consensus statements and discussions of clinical evidence This is a must have reference for researchers and clinicians in translational research *Fossil Energy Update*, 1984

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