

10.1 Periodic Trends



10.1 PERIODIC TRENDS FROM CONCEPTS

Section 10.1 (continued)

In your notebook, read aloud the modern periodic table.

Copy the information in the box on the left below from the periodic table to complete the table on the right.

_____	1
_____	2
_____	3
_____	4
_____	5
_____	6
_____	7

Atomic Mass	_____
Atomic Number	_____
Electron Configuration	_____
Chemical Name	_____
Chemical Symbol	_____

Put each item in Column A, either the letter of the matching item in Column B.

Column A

- _____ 16. It increases in the periodic table.
- _____ 17. It decreases in the periodic table.
- _____ 18. Group 1 elements.
- _____ 19. Elements that are shiny and conduct electricity.
- _____ 20. Group 17 elements.

Column B

- a. metals
- b. group
- c. period
- d. representative elements
- e. transition elements

In the space on the left, write *yes* if the statement is true, if the statement is false, change the underlined word or phrase to make it true.

- _____ 21. There are two representative elements of elements.
- _____ 22. Most elements are metals in the periodic table.
- _____ 23. Group 1 elements are highly reactive metals in the periodic table.
- _____ 24. Group 17 elements are the halogens in the periodic table.
- _____ 25. Group 1 elements are highly reactive metals in the periodic table.
- _____ 26. Group 17 elements are highly reactive nonmetals in the periodic table.
- _____ 27. Group 1 elements are highly reactive metals in the periodic table.
- _____ 28. Elements that are shiny and conduct electricity are metals.

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Doctor Chemistry



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immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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