

CHEMISTRY

2c(ii)

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

where:

 V_1 = Initial Volume = 150 cm^3 T_1 = Initial Temp in kelvin = $21^\circ\text{C} + 273 = 294\text{K}$ T_2 = Final Temp in kelvin = $45^\circ\text{C} + 273 = 318\text{K}$ $V_2 = ?$

$$V_2 = V_1 \times \frac{T_2}{T_1}$$

$$V_2 = 150 \text{ cm}^3 \times \frac{318\text{K}}{294\text{K}}$$

$$V_2 = 150 \text{ cm}^3 \times 1.0816$$

$$V_2 = 162.24 \text{ cm}^3$$

2d.

I Sodium oxide:



II Phosphorus (III) oxide:



III Chlorine (VII) oxide



IV Magnesium oxide



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Richard H. Langley, John Moore



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