Chemistry

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(Chapter 7) (Alcohols Phenols and Ethers) (Exercises)

(Class 12)

Question 7.10:

Write chemical reaction for the preparation of phenol from chlorobenzene.

__Answer 7.10:

Chlorobenzene is fused with NaOH (at 623 K and 320 atm pressure) to produce sodium phenoxide, which gives phenol on acidification.

Question 7.11:

Write the mechanism of hydration of ethene to yield ethanol.

__Answer 7.11:

The mechanism of hydration of ethene to form ethanol involves three steps.

Step 1:

Protonation of ethene to form carbocation by electrophilic attack of H₃O+:

$$H > C = C \le \frac{H}{H} + H = \frac{\ddot{Q}_{*} - H}{H} \longrightarrow H = \frac{\ddot{Q}}{H} - \frac{C}{4} \le \frac{H}{H} + H^{2} \ddot{Q}$$

$$H^{2}O + H_{*} \longrightarrow H^{2}O.$$

Step 2:

Nucleophilic attack of water on carbocation:

$$H - \bigcap_{H} - \bigcap_{H} \bigcap_{H} + H \bigcap_{H} \longrightarrow H - \bigcap_{H} - \bigcap_{H$$

Step 3:

Deprotonation to form ethanol:

Question 7.12:

You are given benzene, conc. H_2SO_4 and NaOH. Write the equations for the preparation of phenol using these reagents.

__Answer 7.12:

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AN Whitehead

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