

RATE LAW EQUATION PROBLEM SET

Part 1: Rate Law Equation Practice

1. For a reaction where the rate equation is $r = k[\text{NH}_4^+_{(\text{aq})}][\text{NO}_2^-_{(\text{aq})}]$,
- calculate k at temperature T_1 , if the rate, r , is $2.40 \times 10^{-7} \text{ mol}/(\text{L} \cdot \text{s})$ when $[\text{NH}_4^+_{(\text{aq})}]$ is 0.200 mol/L and $[\text{NO}_2^-_{(\text{aq})}]$ is 0.00500 mol/L .
 - calculate r at temperature T_2 , if the rate constant, k , is $3.20 \times 10^{-4} \text{ L}/(\text{mol} \cdot \text{s})$ when $[\text{NH}_4^+_{(\text{aq})}]$ is 0.100 mol/L and $[\text{NO}_2^-_{(\text{aq})}]$ is 0.0150 mol/L .
2. A series of experiments is performed for the system
- $$2\text{A} + 3\text{B} + \text{C} \longrightarrow \text{D} + 2\text{E}$$
- When the initial concentration of A is doubled, the rate increases by a factor of 4.
 - When the initial concentration of B is doubled, the rate is doubled.
 - When the initial concentration of C is doubled, there is no effect on rate.
- What is the order of reaction with respect to each of the reactants?
 - Write an expression for the rate equation.

3. The experimental observations in **Table 4** are obtained for the reaction:
- $$2\text{A} + \text{B} + 2\text{C} \rightarrow 3\text{X}$$
- What is the order of reaction with respect to each of the reactants?
 - Write an expression for the rate equation.
 - Calculate a value for the rate constant.
 - Calculate the rate of production of X when $[\text{A}] = [\text{B}] = [\text{C}] = 0.40 \text{ mol/L}$.

Table 4. Observations on the Rate of Production of X

Trial	Initial [A] (mol/L)	Initial [B] (mol/L)	Initial [C] (mol/L)	Rate of production of X (mol/(L·s))
1	0.10	0.10	0.10	3.0×10^{-4}
2	0.20	0.10	0.10	1.2×10^{-3}
3	0.10	0.30	0.10	3.0×10^{-4}
4	0.20	0.10	0.20	2.4×10^{-3}

Part 2: More Rate Law Expressions

1. Consider the data for hydrogen concentration $[\text{H}_2]$, iodine concentration $[\text{I}_2]$ and rate of reaction (moles per litre per second or $\text{mol} \cdot \text{L}^{-1}/\text{s}$) for this reaction:



Trial	$[\text{H}_2]$ (mol/L)	$[\text{I}_2]$ (mol/L)	Rate ($\text{mol} \cdot \text{L}^{-1}/\text{s}$)
1	0.01	0.05	0.04
2	0.02	0.05	0.08
3	0.03	0.05	0.12
4	0.05	0.01	0.02
5	0.05	0.02	0.16
6	0.05	0.03	0.54

- Determine the rate law expression for this reaction.
- What is the overall reaction order?
- What would happen to the rate of the reaction if the concentration of both reactants was doubled?

Rate Law Practice Problems With Answers

R Sanford



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