

TOPIC 5 RATE OF REACTION

ILLUSTRATIVE EXPERIMENTS

- Keep demonstration simple and eye catching
- Do not overlook simple demonstrations. Even trivial activities can spark all sorts of interesting ideas.
- Involve your audience as much as possible; use your students as assistants
- Rehearse the demonstration before hand.
- Demonstrations are not replacements for practical work, but they make good extensions to a topic and help prompt discussion.
- The most important part of a good demonstration is that the students enjoy it.

Changing the speed of reactions. What makes a reaction go faster?

1. Particle size

Perhaps the size of the particles which are reacting has an effect on the speed. The reaction between marble chips and acid can be used to investigate the effect of particle size on the speed of reaction.

Two experiments are set up. The first uses dilute acid and large marble chips. The second has the same concentration of acid and the same mass of marble, however now the marble is in the form of small chips.

In both cases the speed of the reaction is found by measuring the mass of the flasks every minute. The flasks become lighter as the carbon dioxide gas escapes. A simple subtraction gives the mass of carbon dioxide produced.



2. Catalyst

There is a way to speed up a reaction without increasing the temperature or the concentration of the reactants. That is, by adding a substance called a catalyst. Hydrogen peroxide solution is a liquid which decomposes slowly to give oxygen and water, as shown in the equation:



Normally this reaction is so slow that you cannot see any signs of it taking place. But when manganese dioxide is added, bubbles of oxygen are produced very quickly. We can say the manganese dioxide is a catalyst for the decomposition of hydrogen peroxide.

Chemists have discovered many different catalysts. However, all catalysts:

- Speed up some reactions
- Are not used up during a reaction.

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