

REPORT SHEET

Rates of Chemical Reaction

Name _____

Date _____

Section _____

Partner's Name _____

Instructor _____

(1.5 pts. total for Tables 1-3)

Table 1, Experiment #1

Initial moles of $S_2O_8^{2-}$		Cumulative time recorded from stopwatch		Cumulative time converted to seconds (s)	Mole of $S_2O_8^{2-}$ Consumed	Mole of $S_2O_8^{2-}$ Remaining	Concentration of $S_2O_8^{2-}$ Remaining (mol/L)
Aliquot #		(min)	(sec)				
1		1 min	46 sec	106	$1.0 \times 10^{-4} \text{ mol}$	0.025 mol	0.046 M
2		3	24 sec	204	$2.0 \times 10^{-4} \text{ mol}$	0.023 mol	0.043 M
3		4	29 sec	269	$3.0 \times 10^{-4} \text{ mol}$	0.021 mol	0.041 M
4		5	48 sec	348	$4.0 \times 10^{-4} \text{ mol}$	0.020 mol	0.038 M
5		7	03	423	$5.0 \times 10^{-4} \text{ mol}$	0.019 mol	0.035 M

Table 2, Experiment #2

Initial moles of $S_2O_8^{2-}$		Cumulative time recorded from stopwatch		Cumulative time converted to seconds (s)	Mole of $S_2O_8^{2-}$ Consumed	Mole of $S_2O_8^{2-}$ Remaining	Concentration of $S_2O_8^{2-}$ Remaining (mol/L)
Aliquot #		(min)	(sec)				
1		2	46	165	$1.0 \times 10^{-4} \text{ mol}$	0.023 mol	0.046 M
2		5	19	319	$2.0 \times 10^{-4} \text{ mol}$	0.022 mol	0.043 M
3		8	33	513	$3.0 \times 10^{-4} \text{ mol}$	0.021 mol	0.041 M
4		11	11	671	$4.0 \times 10^{-4} \text{ mol}$	0.020 mol	0.038 M
5		13	17	797	$5.0 \times 10^{-4} \text{ mol}$	0.019 mol	0.035 M

Table 3, Experiment #3

Initial moles of $S_2O_8^{2-}$		Cumulative time recorded from stopwatch		Cumulative time converted to seconds (s)	Mole of $S_2O_8^{2-}$ Consumed	Mole of $S_2O_8^{2-}$ Remaining	Concentration of $S_2O_8^{2-}$ Remaining (mol/L)
Aliquot #		(min)	(sec)				
1		1	20	80	$1.0 \times 10^{-4} \text{ mol}$	0.024 mol	0.034 M
2		2	30	150	$2.0 \times 10^{-4} \text{ mol}$	0.024 mol	0.035 M
3		3	43	223	$3.0 \times 10^{-4} \text{ mol}$	0.025 mol	0.037 M
4		3	59	233	$4.0 \times 10^{-4} \text{ mol}$	0.024 mol	0.033 M
5		5	06	306	$5.0 \times 10^{-4} \text{ mol}$	0.023 mol	0.030 M

Rates Of Chemical Reactions Lab

Christian Drosten



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Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations *Environmental & Water*

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emphasis on handling data is supported by the Data Analysis section **EduGorilla's CBSE Class 12th Chemistry Lab Manual | 2024 Edition | A Well Illustrated** EduGorilla Prep Experts, *General Catalog* Georgia Institute of

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Carnicom, 1968 The values used by a number of investigators for the rate constants of high temperature greater than or equal to 1000 K homogeneous gaseous reactions involving species of the elements nitrogen oxygen carbon and sodium have

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