

## Worksheet 1-1 - Measuring Reaction Rates

Name - \_\_\_\_\_  
Due Date - \_\_\_\_\_

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

1. A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:



A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

$$\begin{array}{r} 0.016\text{g} \\ - 0.006\text{g} \\ \hline 0.010\text{g} \end{array}$$

- a) Calculate the *Rate of Reaction* in grams of Zn consumed per second.

$$\text{rate} = \frac{\Delta g}{\Delta \text{time}} = \frac{0.010\text{g}}{20\text{sec}}$$

Answer  $5.0 \times 10^{-4}$

- b) Calculate the *Rate of Reaction* in moles of Zn consumed per second.

$$5.0 \times 10^{-4} \frac{\text{g}}{\text{sec}} \times \frac{1\text{mol}}{65.4\text{g}} =$$

Answer  $7.6 \times 10^{-6} \frac{\text{mol}}{\text{sec}}$

- c) Write out the complete ionic equation for the reaction.



- d) What will happen to the  $[\text{H}^+]$  as the reaction proceeds?

decrease

- e) What will happen to the  $[\text{Cl}^-]$  as the reaction proceeds?

nothing, it is a spectator

# Rate Of Reaction 1 Answers Key

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TABLE OF CONTENT

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## **Table of Contents Rate Of Reaction 1 Answers Key**

1. Understanding the eBook Rate Of Reaction 1 Answers Key
  - The Rise of Digital Reading Rate Of Reaction 1 Answers Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Rate Of Reaction 1 Answers Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Rate Of Reaction 1 Answers Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Rate Of Reaction 1 Answers Key
  - Personalized Recommendations

- Rate Of Reaction 1 Answers Key User Reviews and Ratings
- Rate Of Reaction 1 Answers Key and Bestseller Lists
- 5. Accessing Rate Of Reaction 1 Answers Key Free and Paid eBooks
  - Rate Of Reaction 1 Answers Key Public Domain eBooks
  - Rate Of Reaction 1 Answers Key eBook Subscription Services
  - Rate Of Reaction 1 Answers Key Budget-Friendly Options
- 6. Navigating Rate Of Reaction 1 Answers Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Rate Of Reaction 1 Answers Key Compatibility with Devices
  - Rate Of Reaction 1 Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Rate Of Reaction 1 Answers Key
  - Highlighting and Note-Taking Rate Of Reaction 1 Answers Key
  - Interactive Elements Rate Of Reaction 1 Answers Key
- 8. Staying Engaged with Rate Of Reaction 1 Answers Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Rate Of Reaction 1 Answers Key
- 9. Balancing eBooks and Physical Books Rate Of Reaction 1 Answers Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Rate Of Reaction 1 Answers Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Rate Of Reaction 1 Answers Key
  - Setting Reading Goals Rate Of Reaction 1 Answers Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Rate Of Reaction 1 Answers Key
  - Fact-Checking eBook Content of Rate Of Reaction 1 Answers Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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