

Chemical Kinetics

RATE LAWS

1. Consider the reaction: $2 \text{NO(g)} + \text{O}_2\text{(g)} \rightarrow 2 \text{NO}_2\text{(g)}$

The following data were obtained from three experiments using the method of initial rates:

	Initial [NO] mol L ⁻¹	Initial [O ₂] mol L ⁻¹	Initial rate NO mol L ⁻¹ s ⁻¹
Experiment 1	0.010	0.010	2.5×10^{-5}
Experiment 2	0.020	0.010	1.0×10^{-4}
Experiment 3	0.010	0.020	5.0×10^{-5}

- Determine the order of the reaction for each reactant. **NO 2nd O₂ 1st**
- Write the rate equation for the reaction. **rate = k [NO]² [O₂]**
- Calculate the rate constant. **k = 25 L² mol⁻² s⁻¹**
- Calculate the rate (in mol L⁻¹s⁻¹) at the instant when [NO] = 0.015 mol L⁻¹ and [O₂] = 0.0050 mol L⁻¹. **rate = 2.8 x 10⁻⁵ mol L⁻¹ s⁻¹**
- At the instant when NO is reacting at the rate 1.0 x 10⁻⁴ mol L⁻¹s⁻¹, what is the rate at which O₂ is reacting and NO₂ is forming?
NO: -5.0 x 10⁻⁵ mol L⁻¹s⁻¹ ; NO₂: +1.0 x 10⁻⁴ mol L⁻¹s⁻¹

2. The reaction $2 \text{NO(g)} + 2 \text{H}_2\text{(g)} \rightarrow \text{N}_2\text{(g)} + 2 \text{H}_2\text{O(g)}$ was studied at 904 °C, and the data in the table were collected.

	Initial [NO] mol L ⁻¹	Initial [H ₂] mol L ⁻¹	Initial rate N ₂ mol L ⁻¹ s ⁻¹
Experiment 1	0.420	0.122	0.136
Experiment 2	0.210	0.122	0.0339
Experiment 3	0.210	0.244	0.0678
Experiment 4	0.105	0.488	0.0339

- Determine the order of the reaction for each reactant. **NO 2nd H₂ 1st**
- Write the rate equation for the reaction. **rate = k [NO]² [H₂]**
- Calculate the rate constant at 904 °C. **k = 6.32 L² mol⁻² s⁻¹**
- Find the rate of appearance of N₂ at the instant when [NO] = 0.350 M and [H₂] = 0.205 M. **rate = 0.159 mol L⁻¹s⁻¹**

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JG Myers



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Table of Contents Rates Of Reaction 18 1 Practice Problems Worksheey

1. Understanding the eBook Rates Of Reaction 18 1 Practice Problems Worksheey
 - The Rise of Digital Reading Rates Of Reaction 18 1 Practice Problems Worksheey
 - Advantages of eBooks Over Traditional Books
2. Identifying Rates Of Reaction 18 1 Practice Problems Worksheey
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Rates Of Reaction 18 1 Practice Problems Worksheey
 - User-Friendly Interface
4. Exploring eBook Recommendations from Rates Of Reaction 18 1 Practice Problems Worksheey
 - Personalized Recommendations
 - Rates Of Reaction 18 1 Practice Problems Worksheey User Reviews and Ratings
 - Rates Of Reaction 18 1 Practice Problems Worksheey and Bestseller Lists
5. Accessing Rates Of Reaction 18 1 Practice Problems Worksheey Free and Paid eBooks
 - Rates Of Reaction 18 1 Practice Problems Worksheey Public Domain eBooks
 - Rates Of Reaction 18 1 Practice Problems Worksheey eBook Subscription Services
 - Rates Of Reaction 18 1 Practice Problems Worksheey Budget-Friendly Options

6. Navigating Rates Of Reaction 18 1 Practice Problems Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Rates Of Reaction 18 1 Practice Problems Worksheet Compatibility with Devices
 - Rates Of Reaction 18 1 Practice Problems Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Rates Of Reaction 18 1 Practice Problems Worksheet
 - Highlighting and Note-Taking Rates Of Reaction 18 1 Practice Problems Worksheet
 - Interactive Elements Rates Of Reaction 18 1 Practice Problems Worksheet
8. Staying Engaged with Rates Of Reaction 18 1 Practice Problems Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Rates Of Reaction 18 1 Practice Problems Worksheet
9. Balancing eBooks and Physical Books Rates Of Reaction 18 1 Practice Problems Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Rates Of Reaction 18 1 Practice Problems Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Rates Of Reaction 18 1 Practice Problems Worksheet
 - Setting Reading Goals Rates Of Reaction 18 1 Practice Problems Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Rates Of Reaction 18 1 Practice Problems Worksheet
 - Fact-Checking eBook Content of Rates Of Reaction 18 1 Practice Problems Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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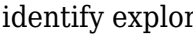
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