

Writing Balanced Precipitation Reactions

In this problem, you will go into the virtual laboratory and perform a series of precipitation reactions using Ag^+ , Pb^{2+} , and Sb^{3+} . After observing the reactions, you will write the net ionic equations representing these reactions and then balance them.

1. Start *Virtual ChemLab*, select *Reactions and Stoichiometry*, and then select *Writing Balanced Precipitation Reactions* from the list of assignments. The lab will open in the Inorganic laboratory.
2. Drag a test tube from the box and place it on the metal test tube stand. You can then click on the bottle of Ag^+ ion solution (it is AgNO_3) on the shelf to add it to the test tube. Then drag the test tube to the blue test tube rack. Repeat for Pb^{2+} ($\text{Pb}(\text{NO}_3)_2$) and Sb^{3+} ($\text{Sb}(\text{OH})_3$).
3. Moving your mouse over each test tube in the test-tube rack will allow you to identify it on the TV monitor. You can also click on the label at the top of each test tube in order to label the test tubes. Place the test tube containing the Ag^+ solution in the metal test tube stand. You will be able to observe what happens in the window at the top left. Click on the Na_2CO_3 reagent bottle to add it to the test tube in the stand.

What color is the precipitate? Write a correct balanced net ionic equation for the reaction.

Brown

When you have finished writing the equation and making your observations, place the test tube containing the precipitate in the red disposal bucket.

4. Place the test tube containing the Pb^{2+} solution in the metal test tube stand. Click on the NaCl reagent bottle to add it to the test tube in the stand.

What color is the precipitate? Write a correct balanced net ionic equation for this reaction.

White

When you have finished writing the equation and making your observations, place the test tube containing the precipitate in the red disposal bucket.

5. Place the test tube containing the Sb^{3+} solution in the metal test tube stand. Note that the antimony is not soluble and you will have a double displacement reaction. Click on the Na_2S reagent bottle to add it to the test tube in the stand.

What color is the precipitate? Write a correct balanced net ionic equation for the reaction.

Orange

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