

# Basic Chemistry Study Guide 2

## 1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's  $\delta^+$  region is attracted to a  $\delta^-$  region in another molecule.
3. Form between water molecules (Where are the  $\delta^+$  and  $\delta^-$  regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

## 2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

## 3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

# Upper Secondary Chemistry Study Guide

**Xiang Xie**



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## **Table of Contents Upper Secondary Chemistry Study Guide**

1. Understanding the eBook Upper Secondary Chemistry Study Guide
  - The Rise of Digital Reading Upper Secondary Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Upper Secondary Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Upper Secondary Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Upper Secondary Chemistry Study Guide
  - Personalized Recommendations
  - Upper Secondary Chemistry Study Guide User Reviews and Ratings
  - Upper Secondary Chemistry Study Guide and Bestseller Lists

5. Accessing Upper Secondary Chemistry Study Guide Free and Paid eBooks
  - Upper Secondary Chemistry Study Guide Public Domain eBooks
  - Upper Secondary Chemistry Study Guide eBook Subscription Services
  - Upper Secondary Chemistry Study Guide Budget-Friendly Options
6. Navigating Upper Secondary Chemistry Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Upper Secondary Chemistry Study Guide Compatibility with Devices
  - Upper Secondary Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Upper Secondary Chemistry Study Guide
  - Highlighting and Note-Taking Upper Secondary Chemistry Study Guide
  - Interactive Elements Upper Secondary Chemistry Study Guide
8. Staying Engaged with Upper Secondary Chemistry Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Upper Secondary Chemistry Study Guide
9. Balancing eBooks and Physical Books Upper Secondary Chemistry Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Upper Secondary Chemistry Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Upper Secondary Chemistry Study Guide
  - Setting Reading Goals Upper Secondary Chemistry Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Upper Secondary Chemistry Study Guide
  - Fact-Checking eBook Content of Upper Secondary Chemistry Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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