

## AP Questions: Thermochemistry

1984 B

Substance	Standard Heat of Formation, $\Delta H_f^\circ$ in kJ/mol
C(s)	0.00
CO <sub>2</sub> (g)	-393.5
H <sub>2</sub> (g)	0.00
H <sub>2</sub> O(l)	-285.85
O <sub>2</sub> (g)	0.00
C <sub>3</sub> H <sub>7</sub> COOH(l)	?

The enthalpy change for the combustion of butyric acid at 25°C,  $\Delta H_{\text{comb}}$ , is -2,183.5 kilojoules per mole. The combustion reaction is



- From the above data, calculate the standard heat of formation,  $\Delta H_f^\circ$ , for butyric acid.
- Write a correctly balanced equation for the formation of butyric acid from its elements.

1988 B

Substance	Enthalpy of Combustion, $\Delta H$ (kilojoules/mole)
C(s)	-393.5
H <sub>2</sub> (g)	-285.8
C <sub>2</sub> H <sub>5</sub> OH(l)	-1366.7
H <sub>2</sub> O(l)	--

- Write a separate, balanced chemical equation for the combustion of each of the following: C(s), H<sub>2</sub>(g), and C<sub>2</sub>H<sub>5</sub>OH(l). Consider the only products to be CO<sub>2</sub> and/or H<sub>2</sub>O(l).
- In principle, ethanol can be prepared by the following reaction:



Calculate the standard enthalpy change,  $\Delta H$ , for the preparation of ethanol, as shown in the reaction above.

1988 D

An experiment is to be performed to determine the standard molar enthalpy of neutralization of a strong acid by a strong base. Standard school laboratory equipment and a supply of standardized 1.00 molar HCl and standardized 1.00 molar NaOH are available.

- What equipment would be needed?
- What measurements should be taken?
- Without performing calculations, describe how the resulting data should be used to obtain the standard molar enthalpy of neutralization.
- When a class of students performed this experiment, the average of the results was -55.0 kilojoules per mole. The accepted value for the standard molar enthalpy of neutralization of a strong acid by a strong base is -57.7 kilojoules per mole. Propose two likely sources of experimental error that could account for the result obtained by the class.

1995 B

Propane, C<sub>3</sub>H<sub>8</sub>, is a hydrocarbon that is commonly used as fuel for cooking.

- Write a balanced equation for the complete combustion of propane gas, which yields CO<sub>2</sub>(g) and H<sub>2</sub>O(l).
- Calculate the volume of air at 30°C and 1.00 atmosphere that is needed to burn completely 10.0 grams of propane. Assume that air is 21.0 percent O<sub>2</sub> by volume.
- The heat of combustion of propane is -2,220.1 kJ/mol. Calculate the heat of formation,  $\Delta H_f^\circ$ , of propane given that  $\Delta H_f^\circ$  of H<sub>2</sub>O(l) = -285.3 kJ/mol and  $\Delta H_f^\circ$  of CO<sub>2</sub>(g) = -393.5 kJ/mol.
- Assuming that all of the heat evolved in burning 30.0 grams of propane is transferred to 8.00 kilograms of water (specific heat = 4.18 J/g°C), calculate the increase in temperature of water.

# Objectives Questions On Thermochemistry

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## Engineering Chemistry I (for BPUT) Patrizia Bocchetta, 2024

*General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions. It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course. It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material.

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Understanding General Chemistry Atef Korchef, 2022-03-07 Understanding General Chemistry details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning the book features main objectives at the beginning of each chapter, get smart sections and check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs exercises and problems with the corresponding answers and explanations are readily available. Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self evaluate their skill level. Includes powerful methods, easy steps, simple and accurate interpretations and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry and physical chemistry. This introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students. **Iit Objective Chemistry** Arun Syamal, 2008-08

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