

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reactants. These reactions are called reversible reactions, as the reaction can proceed in either direction. Once the equilibrium is reached the forward and reverse reactions continue at equal rates and the concentrations of reactants and products remain constant. The importance of understanding that equilibrium is a dynamic state is that it allows us to predict the behavior of chemical systems. For example, if a reaction is at equilibrium, then the forward and reverse reactions continue to occur over time, but there is no net change in the concentrations of reactants and products. Equilibrium conditions are reached only at constant pressure, or concentration, or volume, or temperature. Le Chatelier's Principle states that applying stress changes the equilibrium position of an equilibrium system in a direction that relieves the stress applied. Stress can be a change in concentration, pressure, or temperature. For the change in concentration, if a reactant is added, the equilibrium will shift right to consume the added reactant. If a product is added, the equilibrium will shift left to consume the added product. If the stress is a change in temperature, then the equilibrium will shift in a direction that absorbs the heat added. For the change in pressure, if the pressure is increased, the equilibrium will shift in a direction that decreases the number of moles of gas. For example, if the reaction is at equilibrium, then the reaction will shift in a direction that decreases the number of moles of gas if the pressure is increased.



Adding NaSCN increases $\text{FeSCN}^{2+}(\text{aq})$ and shifts the $\text{Fe}^{3+}(\text{aq})$ to the right to consume the added SCN}^{-}. Adding the reactant increases the concentration of the reactant and the equilibrium shifts in a direction that consumes the added reactant.

Procedure (This Group):

- | | |
|--|---|
| 1. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 2. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 3. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 4. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 5. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 6. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 7. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 8. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 9. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 10. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 11. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 12. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 13. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 14. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 15. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 16. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 17. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 18. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 19. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 20. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 21. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 22. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 23. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 24. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 25. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 26. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 27. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 28. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 29. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 30. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 31. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 32. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 33. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 34. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 35. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 36. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 37. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 38. Prepare 100 mL of 0.1 M SCN^{-} solution. |
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| 47. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 48. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 49. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 50. Prepare 100 mL of 0.1 M SCN^{-} solution. |
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| 97. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 98. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 99. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 100. Prepare 100 mL of 0.1 M SCN^{-} solution. |

Observing Chemical Equilibrium Lab 46

WJ Hussar



Observing Chemical Equilibrium Lab 46:

Unveiling the Magic of Words: A Report on "**Observing Chemical Equilibrium Lab 46**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Observing Chemical Equilibrium Lab 46**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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