

## Practice: pages 20–25

- 1 Answer is B. From experiments 1 and 3 the order with respect to  $\text{Cl}_2$  is 1 (the rate quadruples as the concentration doubles. Therefore, in an experiment using  $0.80 \text{ mol dm}^{-3}$  of CO and  $0.20 \text{ mol dm}^{-3}$  of  $\text{Cl}_2$ , the rate will be halved from the  $6.4 \times 10^{-24} \text{ mol dm}^{-3} \text{ min}^{-1}$  obtained in experiment 1. (There is no need to work out the order with respect to CO.) [1]
- 2 Answer is B. The concentration halves every 120 s. So, the sequence of half-lives is  $2.4 \rightarrow 1.2 \rightarrow 0.6 \rightarrow 0.3 \rightarrow 0.15$ ; that is four half-lives which is  $4 \times 120 \text{ s}$ . [1]
- 3 Answer is C. For a zero order reaction, the rate of decomposition is a constant. [1]
- 4 Answer is A. the decomposition is  $\text{Rate} = k[\text{X}]^2$  [1]  
From the data given,  $0.016 = k[1]^2$  [1]  
Therefore,  $k = 0.016 \text{ s}^{-1}$ . If  $[\text{X}] = 0.25 \text{ mol dm}^{-3}$ ,  $\text{rate} = 0.016[0.25]^2$  [1]  
 $= 0.001 \text{ dm}^3 \text{ mol}^{-1} \text{ s}^{-1}$  [1]
- 5 Answer is B. The overall order is 2, so the units are  $\frac{1}{\text{concentration} \times \text{s}}$ . [1]
- 6 Answer is D.  $\text{I}_2$  has an order of zero and is not involved in the rate-determining step. Therefore, it can be ignored. In D,  $[0.1][2]$  has a value that is greater than  $[0.1][1]$ . [1]
- 7 Answer is C. The slow step + the fast step must give the overall equation. I must, therefore, be removed in the fast step in a reaction with G, and some J and K must be formed. [1]
- 8 Answer is B. The rate equation is  $\text{Rate} = k[\text{halide}]$  so statements 1 and 2 are correct. Raising the temperature will increase the rate of reaction by allowing more molecules to exceed the activation energy, but the activation energy is fixed and will not change. [1]
- 9 Answer is D. Increasing the amount of ozone present does not change the order with respect to ozone. A large excess is not necessary to ensure enough ozone is present. The purpose of a large excess is to make sure that during the reaction its concentration is virtually constant. Therefore, only statement 3 is correct. [1]
- 10 Answer is A. Experiments 2 and 3 give the order with respect to X as zero. Experiments 1 and 2 give the order with respect to Y as 2. So statements 1 and 2 are correct. The rate constant is calculated from a set of results, for example, experiment 1 gives  $0.08 = k[0.10]^2$ .  
So  $k = 8.0 \text{ dm}^3 \text{ mmol}^{-1} \text{ s}^{-1}$  and statement 3 is, therefore, correct. [1]
- 11 a)  $2\text{N}_2\text{O}_5(\text{g}) \rightarrow 4\text{NO}_2(\text{g}) + \text{O}_2(\text{g})$  [1]  
b) i) For every 2 mol of  $\text{N}_2\text{O}_5$  that reacts 4 mol of  $\text{NO}_2$  is formed. Therefore, the rate of formation of  $\text{NO}_2$  is twice as fast as the loss of  $\text{N}_2\text{O}_5$ , i.e.  $7.0 \times 10^{-4} \text{ mol dm}^{-3} \text{ s}^{-1}$ . [1]  
ii) For oxygen, the rate of formation will be  $\frac{1}{2}$  the rate of loss of  $\text{N}_2\text{O}_5$ ,  
i.e.  $1.75 \times 10^{-4} \text{ mol dm}^{-3} \text{ s}^{-1}$  [1]

# Reaction Rate Edgenuity Answers

**D. Heidrich**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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**Some Statistical Methods for the Analysis of Chemical Reaction Rate Data** Anna Christine Miller, 1960 Reaction Rate of the Deacon Reaction Robert H. Johnston, 1945 **The Determination of Reaction-rate Constants**, 1972 *Reaction Rate Measurements at the National Criticality Experiments Research Center*, 2014 *The  $^{15}\text{O}(\alpha, \gamma)^{19}\text{Ne}$  Reaction Rate and the Rp Process* Paul Vogt Magnus, 1988 **Approximation of Reaction Rate Models** Hsu-Tung Lee, 1967 Reaction Rate Handbook United States. Defense Nuclear Agency, 1972 *Reaction Rate Prediction for Elementary Reactions in Binary Mixed Solvents Assuming an Additive Solvation of Reactant(s)* Howard Clark Van Woert, 1976 **Reaction Rate Handbook** United States. Defense Nuclear Agency, 1972 The Reaction Path in Chemistry: Current Approaches and Perspectives D. Heidrich, 2013-03-09 The so called reaction path RP with respect to the potential energy or the Gibbs energy free enthalpy is one of the most fundamental concepts in chemistry It significantly helps to display and visualize the results of the complex microscopic processes forming a chemical reaction This concept is an implicit component of conventional transition state theory TST The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination However an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and

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