

8.3. Quantitative Balance Problems: Problems

1. Consider the reaction of the gasification of ammonia, (8) in Section 8.1.



It is assumed to proceed at a constant temperature of 200°C and the inlet NH_3 gas stream enters at the inlet gas velocity of 1.5 km/s. Estimate:

- (a) The unsteady-state equations for the concentrations of nitrogen and hydrogen gases, as a function of catalyst weight in the reactor. Assume



- as follows. Write the mass balance equations for these reactions.

- (b) The initial concentrations of the gases entering the reactor are 10% ammonia, 10% carbon monoxide, and oxygen gas was purchased in 20% ammonia. Calculate the rate of reaction of ammonia based on the oxygen gas.

- (c) These reactions are believed to be gas, not solid, catalyzed reactions. Based on the steps in above mechanism, derive an overall rate-law for the reaction of ammonia gas to nitrogen.

- (d) The reaction is gasification of a mixture of gases.

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O.P. Gupta



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