

Equilibrium revision

Questions

Q.1

Chlorine gas (Cl_2) is bubbled through a solution of potassium bromide (KBr) and the reaction is shown in the equation below.



The equilibrium constant (K) for this reaction is 4.2×10^{22}

- a. Give the expression for equilibrium constant

$$K = \frac{[\text{Br}_2] [\text{Cl}^-]^2}{[\text{Cl}_2] [\text{Br}^-]^2}$$

- b. What information about the equilibrium can be deduced from the value of K.
 $K=4.2 \times 10^{22}$ which is very high. This indicates that at equilibrium products are favoured greatly or reaction has gone to almost completion
- c. State 3 factors that will shift the equilibrium to the right.
Increase conc of $\text{Br}^-(\text{aq})$, increase conc of $\text{Cl}_2(\text{g})$, decrease temp

Q.2

What would happen to the position of equilibrium (which direction will it shift) when the following changes are made to the reaction below?



- a Sulfur dioxide is added to the system **Left**
- b Sulfur trioxide is removed from the system **Left**
- c Oxygen is added to the system **Left**

Q.3

The reaction between carbon and water can be represented by the following equilibrium equation:



- (a) Write an expression for the equilibrium constant for this reaction
 $K = \frac{[\text{H}_2] [\text{CO}]}{[\text{H}_2\text{O}]}$
- (b) What would be the effect on the equilibrium
- i) if you increased the concentration of $\text{CO}(\text{g})$? **No effect**
- ii) if you increased the temperature of the system? **K will increase**

Reactions Rates And Equilibrium Section Review Answers

Leo J. Malone



Reactions Rates And Equilibrium Section Review Answers:

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