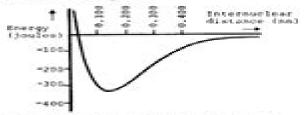
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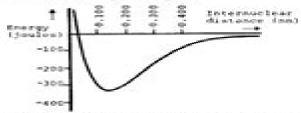
# 9 • Bonding & Molecular Structure

- A bond forms when two orbitals allowing electrons to be close to two nuclei at the same time. When two H atoms bond, they overlap electrons in their \_\_\_\_\_ orbitals.
- As two H atoms approach each other, the potential energy \_\_\_\_\_\_ because of the electron-proton attractions. The potential energy \_\_\_\_\_ because of electron-electron and proton-proton repulsions.
- 3. Consider this diagram:



Two separate atoms have an energy of \_\_\_J.
The equilibrium bond distance is \_\_\_ nm.
The bond energy is \_\_\_\_ Joules which is the amount of energy you need to add to break the bond (i.e. separate the two atoms).

4. Consider this diagram again:



If two He atoms approach, their bond distance is 0.500 nm and bond energy is 100 J. Sketch the energy curve for He2. The two atoms do not get close enough for the electrons to be close to both nuclei at the same time... no bond forms.

We generalize: two \_\_\_\_\_\_ orbitals
CAN make a bond.
Two \_\_\_\_\_ orbitals cannot make a bond.

#### LECTURE NOTES1

You should be able quickly to state the number of valence electrons any atom has based on its position on the periodic table. Write it.

P	Ca	F	As
C	Na	к	н
Si	О	В	Mg

Draw the Lewis electron-dot symbol for each of the following elements.

P	Ca	F	As
C	Na	K	н
Si	О	В	Mg

 Some elements have different "ground state" Lewis symbols. Practice drawing them:

Ground State	Bonding State	Ground State	Bonding State
C	C	Ca	Ca
Al	Al	В	В

 You should be able to recognize the type of bond that forms between two atoms based on the atoms involved. Classify each of the following bonds as Ionic or Covalent.

SO KI
NaCl HCl
CaBr. SiQ.
XcE. MgCl
CO NI

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