

- (b) A student knew that calcium hydroxide could be made by adding calcium to water.

The student added 0.00131 mol of calcium to a beaker containing about 100 cm³ of water.
A reaction took place as shown by the equation below.
All the calcium hydroxide formed was soluble.



- (i) Calculate the mass of calcium that the student added.

mass of calcium = g [1]

- (ii) Calculate the volume of hydrogen gas, in dm³, produced in this reaction at room temperature and pressure, RTP.

volume of hydrogen gas = dm³ [1]

- (iii) The student transferred the contents of the beaker to a 250 cm³ volumetric flask and water was added to make the solution up to 250 cm³.

Calculate the concentration, in mol dm⁻³, of **hydroxide ions** in the 250 cm³ solution.

Ocr Chemistry May 2013 F321

Rachel Sandford



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