

CHAPTER 16 REVIEW

Reaction Energy

SECTION 2

SHORT ANSWER Answer the following questions in the space provided.

1. For the following examples, state whether the change in entropy favors the forward or reverse reaction:

- | | |
|-------------------------|--|
| <u>forward reaction</u> | a. $\text{HCl}(l) \rightleftharpoons \text{HCl}(g)$ |
| <u>reverse reaction</u> | b. $\text{C}_6\text{H}_{12}\text{O}_6(aq) \rightleftharpoons \text{C}_6\text{H}_{12}\text{O}_6(s)$ |
| <u>forward reaction</u> | c. $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$ |
| <u>reverse reaction</u> | d. $3\text{C}_2\text{H}_4(g) \rightleftharpoons \text{C}_6\text{H}_{12}(l)$ |

2. (change in free energy) $\Delta G =$ (change in enthalpy) $\Delta H - T$ (change in entropy) ΔS . Write an equation that shows the relationship between enthalpy, ΔH , entropy, ΔS , and free energy, ΔG .

- negative b. For a reaction to occur spontaneously, the sign of ΔG should be ____

3. Consider the following equation: $\text{NH}_3(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{NH}_4^+(aq) + \text{OH}^-(aq) + \text{energy}$

- true a. The enthalpy factor favors the forward reaction. True or False?
reverse reaction b. The sign of $T\Delta S^\circ$ is negative. This means the entropy factor favors the ____

- c. Given that ΔG° for the above reaction in the forward direction is positive, which term is greater in magnitude and therefore predominates, $T\Delta S$ or ΔH ?

T(change in)S is greater than (change in)H; (change in)H is negative, but (change in)G is positive, indicating the reaction is not spontaneous. The randomness of the reactants and the temperature of the reaction overshadow the exothermic factor.

4. Consider the following equation for the vaporization of water:



- | | |
|-------------------------|--|
| <u>endothermic</u> | a. Is the forward reaction exothermic or endothermic? |
| <u>reverse reaction</u> | b. Does the enthalpy factor favor the forward or reverse reaction? |
| <u>forward reaction</u> | c. Does the entropy factor favor the forward or reverse reaction? |

Reaction Energy Section Answer Key

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Table of Contents Reaction Energy Section Answer Key

1. Understanding the eBook Reaction Energy Section Answer Key
 - The Rise of Digital Reading Reaction Energy Section Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Energy Section Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Energy Section Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Energy Section Answer Key
 - Personalized Recommendations
 - Reaction Energy Section Answer Key User Reviews and Ratings
 - Reaction Energy Section Answer Key and Bestseller Lists

5. Accessing Reaction Energy Section Answer Key Free and Paid eBooks
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 - Reaction Energy Section Answer Key eBook Subscription Services
 - Reaction Energy Section Answer Key Budget-Friendly Options
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 - Reaction Energy Section Answer Key Compatibility with Devices
 - Reaction Energy Section Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Energy Section Answer Key
 - Highlighting and Note-Taking Reaction Energy Section Answer Key
 - Interactive Elements Reaction Energy Section Answer Key
8. Staying Engaged with Reaction Energy Section Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Energy Section Answer Key
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Energy Section Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reaction Energy Section Answer Key
 - Setting Reading Goals Reaction Energy Section Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reaction Energy Section Answer Key
 - Fact-Checking eBook Content of Reaction Energy Section Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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