


AQA A Level Chemistry Presentations

LECTURE 36 - RATES OF REACTION



The diagram shows a gas collection apparatus. A conical flask containing a blue liquid (hydrochloric acid) and small pieces of a grey solid (magnesium) is sealed with a stopper. A delivery tube is inserted through the stopper, with its end submerged in the liquid. The tube leads to an inverted measuring cylinder (gas syringe) that is partially filled with a grey liquid. Bubbles of gas are shown entering the cylinder from the tube. Labels point to the 'pieces of Mg' at the bottom of the flask, the 'acid solution', the 'mouth of H₂ gas delivery tube', and the 'H₂ gas collected in gas syringe'.

The reaction between HCl and Mg produces H₂ gas

$$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$$

The volume of hydrogen gas produced is recorded at regular intervals. The rate of reaction can be determined by how fast the gas volume increases.

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