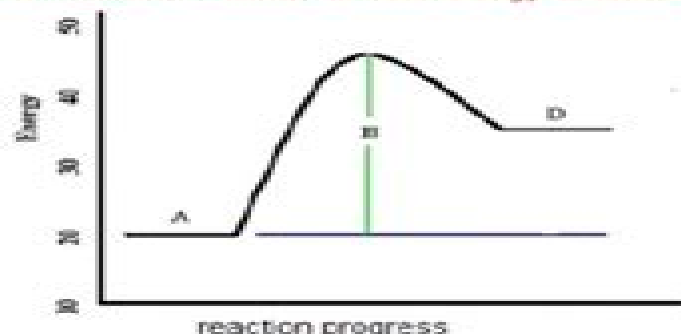


- Suppose two molecules that can react collide. What **3** circumstances need to occur for the molecules to react?  
1. There must be enough energy, 2. They must collide, 3. They must have the correct orientation
- What does the **activation energy** for a chemical reaction mean?  
This is the amount of energy required to make the molecules collide so that a reaction may occur.
- What role does the **reactivity of the reactants** play in determining the rate of a chemical reaction?  
If reactants are not reactive, the reaction may not occur.
- Using the **collision theory**, explain why sugar sprinkled over a Bunsen burner reacts more quickly than a **chunk** of the same solid.  
When the sugar is broken up into individual crystals, there is more surface area and more molecules available to collide and react.
- Use the **collision theory** to explain why **increasing the concentration** of a reactant usually increases the reaction rate.  
Increasing the concentration increases the number of molecules which increases the likelihood that collisions will occur.
- Use the **collision theory** to explain why **increasing the temperature** usually increases the reaction rate.  
Increasing the temperature increases the speed of the molecules. Faster moving molecules collide more.
- Explain how a **catalyst** affects the activation energy for a chemical reaction.  
A catalyst ensure that molecules will collide in the correct orientation, so less energy is needed.
- On** the accompanying energy diagram, label the following items:
  - reactants **A**
  - products **D**
  - activation energy **B**
- Draw and label what this diagram would look like if a **catalyst** was added to the reaction.
- Graph reading
  - How much energy do the reactants have?  
**20 kJ**
  - How much energy do the products have?  
**35 kJ**
  - How much energy is required to activate this un-catalyzed reaction?  
**Approximately 25 kJ.**
  - Will this reaction need to absorb energy or will it release energy to the environment? How do you know?  
**It will absorb energy because the products have more energy than the reactants.**



# Reaction Rates Worksheet With Answers

**EW Minium**



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