- Draw a graph to show the relationship between the number of carbon atoms in each alkane and its boiling point. (Number of carbon atoms will go on the x-axis and boiling point on the y-axis).
- Describe what you see.
- 3. Suggest a reason for what you have observed.
- 4. Why was it enough for us to use 'number of carbon atoms' as a measure of the molecular weight of the molecules?

Click here for the solution^b

2.5.2 Investigation: Determining the density of liquids:

Density is a very important property because it helps us to identify different materials. Every material, depending on the elements that make it up and the arrangement of its atoms, will have a different density. The equation for density is:

$$Density = \frac{Mass}{Volume}$$
(2.1)

Discussion questions:

To calculate the density of liquids and solids, we need to be able to first determine their mass and volume. As a group, think about the following questions:

- How would you determine the mass of a liquid?
- How would you determine the volume of an irregular solid?

Apparatus:

Laboratory mass balance, 10 ml and 100 ml graduated cylinders, thread, distilled water, two different liquids.

Method:

Determine the density of the distilled water and two liquids as follows:

- 1. Measure and record the mass of a 10 ml graduated cyclinder.
- 2. Pour an amount of distilled water into the cylinder.
- Measure and record the combined mass of the water and cylinder.
- 4. Record the volume of distilled water in the cylinder
- 5. Empty, clean and dry the graduated cylinder.
- Repeat the above steps for the other two liquids you have.
- Complete the table below.

Liquid	Mass (g)	Volume (ml)	Density $(g \cdot ml^{-1})$
Distilled water.			
Liquid 1			
Liquid 2			

Table 2.5

2.5.3 Investigation: Determining the density of irregular solids:

Apparatus:

Use the same materials and equipment as before (for the liquids). Also find a number of solids that have an irregular shape.

Method:

Determine the density of irregular solids as follows:

^{*}https://www.floot.org/BP

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