



SOUND INVESTING

UNCOVER FRAUD AND PROTECT
YOUR PORTFOLIO

Kate Mooney and Kerry Marrer

Sound Investing Uncover Fraud And Protect Your Portfolio

Kate Mooney



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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 120.1 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{100.26 \text{ g}}{120.1 \text{ g}} \times 100 = 83.5\%$

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