

## SS 2 CHEMISTRY PRACTICAL 3<sup>RD</sup> TERM EXMINATION

1. B is a solution containing 5.30g of anhydrous sodium trioxocarbonate (IV) per dm<sup>3</sup>. A is a solution of hydrochloric acid of unknown concentration.  
Put A into the burette and titrate 25cm<sup>3</sup> (or 20cm<sup>3</sup>) portions of B with A. Using methyl orange as the indicator. Record the volume of your pipette.  
Tabulate your burette readings, and calculate the average volume of acid used.  
From your results, calculate:
  - a. The concentration of solution B in moles per dm<sup>3</sup>
  - b. The concentration of solution A in moles per dm<sup>3</sup>
  - c. The mass concentration in g per dm<sup>3</sup> of the acid in solution A

The equation of reaction is :



(H = 1; C = 12; O = 16; Na = 23; Cl = 35.5)

2. (a). If you were provided with anhydrous CaCO<sub>3</sub>, spatula and stirrer (i) list three other materials you would require to prepare a standard solution of CaCO<sub>3</sub> (ii) State what you would observe on adding dilute HCl to a portion of the CaCO<sub>3</sub>  
(b). State what would be observed when aqueous ammonia is added to: (i) Litmus paper; (ii) Pb(NO<sub>3</sub>)<sub>2</sub> solution in drops until in excess.

# Sschemistry Practical 2015 3rd Term

**RJ Alexander**



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