

STAAR Science Tutorial 06

TEK 8.5B: Atomic Structure & Chemical Properties

TEK 8.5B: *Identify that protons determine an element's identity and valence electrons determine its chemical properties, including reactivity.*

Valence Electrons & Chemical Properties

- Each element has a unique number of protons in all of its atoms. The **atomic number** of an element is the number of protons in each of its atoms. To identify which element a model of an atom is depicting, count the number of protons in the nucleus and find the element on the periodic table with that atomic number.
- The total number of electrons in an (electrically neutral) atom equals the number of protons. Thus, an atom can also be identified by the total number of electrons shown in the Electron Cloud Model or Bohr Model.
- In the Bohr Model of the atom, electrons fill different electron orbitals (also known as electron shells or energy levels) from the inside-out. The first orbital can only hold 2 electrons. The second orbital can hold 8 electrons, the third can hold up to 18, and the fourth up to 32, but the outermost orbital can never have more than 8 electrons. Only these outermost electrons, known as **valence electrons**, can interact with other atoms in chemical reactions.
- The following chart summarizes the maximum number of electrons that can be in each orbital (if it is not the outermost orbital):

Orbital Number	1	2	3	4	5	6	7
No. of Electrons	2	8	18	32	32	18	8

- In general, elements with the same number of valence electrons, usually found in the same group on the periodic table, have similar chemical properties.
- For the Representative Elements in Groups 1 - 2 and 13 -18, the number of valence electrons is equal to the ones digit of the group number. For example, there is 1 valence electron in Group 1 elements, and 3 valence electrons in Group 13 elements.
- The following chart summarizes the number of valence electrons for the representative elements:

Group/Family Number	1	2	13	14	15	16	17	18
Valance Electrons	1	2	3	4	5	6	7	8

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