



Underlined words and phrases are to be filled in by students on the Note-taking Worksheet.

Section 1 Stability in Bonding

- A. Some elements combine chemically and no longer have the same properties they did before forming a compound.
- B. A(n) chemical formula is composed of symbols and subscripts indicating the number of atoms of an element in a compound.
- C. Atoms form compounds when the compound is more stable than the separate atoms.
1. Noble gases are more chemically stable than other elements because they have a complete outer energy level.
 2. Elements that do not have full outer energy levels are more stable in compounds.
 3. Atoms can lose, gain, or share electrons to get a stable outer energy level.
 4. A(n) chemical bond is the force that holds atoms together in a compound.

DISCUSSION QUESTION:

Why do elements form compounds? *To become more chemically stable by getting a complete outer energy level*

Section 2 Types of Bonds

- A. A(n) ion is a charged particle because it has more or fewer electrons than protons.
1. When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge.
 2. When an atom gains an electron, it becomes a negatively charged ion.
- B. An ionic compound is held together by the ionic bond—the force of attraction between opposite charges of the ions.
1. The result of this bond is a(n) neutral compound.
 2. The sum of the charges on the ions is zero.
- C. Molecules are neutral particles formed as a result of sharing electrons.
1. A covalent bond is the force of attraction between atoms sharing electrons.
 2. Atoms can form double or triple bonds depending on whether they share two or three pairs of electrons.

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