Study Buddy: Energy and Phase Changes

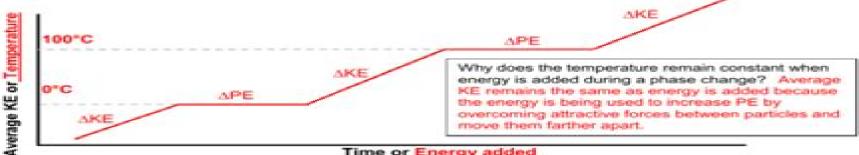
Terms	Description	Picture	Symbol/Unit
Heat	Amount of thermal energy transferred from one system to another	\$	Q/ Joules (J) or calorie (cal)
Total Internal Energy	Describes the total amount of energy in a system. Equals the sum of kinetic and potential energies. Kinetic Energy: due to motion of particles. Potential Energy: due to relative position of particles.	PE or Ephigh	IE/ Joules (J) or calorie (cal)
Endothermic Process	System absorbs energy from surroundings Phase changes: melting, vaporizing, subliming	8	+Q for system
Exothermic Process	System releases heat to surroundings Phase changes: freezing, condensing, depositing		-Q for system
Vapor Pressure	Pressure exerted by the vapor particles over a liquid or solid substance at equilibrium.	STATE OF THE PARTY	http://www.presidence/doile connectingscom/Overmory_/fit presidence/fit/res

How does thermal energy transfer cause phase changes?

As a solid or liquid is heated, the thermal energy is used to increase the kinetic and potential energy of the particles (particles moving faster and farther apart). When a liquid or gas is cooled, the thermal energy transferred to the surroundings reduced the kinetic and potential energies of the particles.



Translating to a Heating or Cooling Curve



Time or Energy added

Label where PE and KE are changing on the heating curve above.

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KJ Lindholm-Leary

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