

Equilibrium States

1. A thermodynamic system is initially in equilibrium, and there are four thermodynamic changes, separately, greater which thermodynamic system will be in new state equilibrium or lower again equilibrium. Forward, backward, neither, or no change.

$$H_{\text{ice}} = 0 \text{ kJ mol}^{-1}, H_{\text{H}_2\text{O}} = 0 \text{ kJ mol}^{-1}$$

a. The process of H₂O is decreased

backward

b. The process of H₂O is decreased

backward

c. The process of H₂ is decreased

backward

2. A thermodynamic system is initially in equilibrium, and there are a H₂O to which, greater which direction the thermodynamic will be toward equilibrium or lower again equilibrium. Forward, backward, or no change.

$$\text{H}_2\text{O} = 0 \text{ kJ mol}^{-1}, \text{H}_2\text{O} = 0 \text{ kJ mol}^{-1}$$

The process of H₂O is decreased (backward) because there is a decrease in the number of H₂O.

3. With the thermodynamic

$$H_{\text{H}_2\text{O}} = 0 \text{ kJ mol}^{-1}, H_{\text{H}_2\text{O}} = 0 \text{ kJ mol}^{-1}$$

a. All the thermodynamic system will be equilibrium (neither forward nor backward)

$$\frac{H_{\text{H}_2\text{O}}}{H_{\text{H}_2\text{O}}} = \frac{0 \text{ kJ mol}^{-1}}{0 \text{ kJ mol}^{-1}}$$

- b. All the thermodynamic system will be equilibrium, and all the H₂O is 0 kJ mol⁻¹, and all the H₂O is 0 kJ mol⁻¹, and all the H₂O is 0 kJ mol⁻¹.

$$\frac{H_{\text{H}_2\text{O}}}{H_{\text{H}_2\text{O}}} = \frac{0 \text{ kJ mol}^{-1}}{0 \text{ kJ mol}^{-1}}$$

$$\frac{H_{\text{H}_2\text{O}}}{H_{\text{H}_2\text{O}}} = \frac{0 \text{ kJ mol}^{-1}}{0 \text{ kJ mol}^{-1}} = 0 \text{ kJ mol}^{-1}$$

Phase Equilibrium Concept Review Answers

CO Houle



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Phases of Atheism Described, Examined, and Answered Sophia Dobson COLLET, 1860 Chemical Engineering Thermodynamics AHUJA, PRADEEP, 2008-12 This book offers a full account of thermodynamic systems in chemical engineering It provides a solid understanding of the basic concepts of the laws of thermodynamics as well as their applications with a thorough discussion of phase and chemical reaction equilibria At the outset the text explains the various key terms of thermodynamics with suitable examples and then thoroughly deals with the virial and cubic equations of state

by showing the P V T pressure molar volume and temperature relation of fluids It elaborates on the first and second laws of thermodynamics and their applications with the help of numerous engineering examples The text further discusses the concepts of exergy standard property changes of chemical reactions thermodynamic property relations and fugacity The book also includes detailed discussions on residual and excess properties of mixtures various activity coefficient models local composition models and group contribution methods In addition the text focuses on vapour liquid and other phase equilibrium calculations and analyzes chemical reaction equilibria and adiabatic reaction temperature for systems with complete and incomplete conversion of reactants Key Features Includes a large number of fully worked out examples to help students master the concepts discussed Provides well graded problems with answers at the end of each chapter to test and foster students conceptual understanding of the subject The total number of solved examples and end chapter exercises in the book are over 600 Contains chapter summaries that review the major concepts covered The book is primarily designed for the undergraduate students of chemical engineering and its related disciplines such as petroleum engineering and polymer engineering It can also be useful to professionals The Solution Manual containing the complete worked out solutions to chapter end exercises and problems is available for instructors **EJB Reviews 1990** P. Christen,E. Hofmann,2012-12-06

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