Molecule Shapes and Inter Molecular Forces

Introduction

Atoms bond to satisfy their need for more electrons. If both atoms have high electro-negativities (are nonmetals), atoms will share electrons to satisfy the Octet Rule – every atom wants 8 electrons to fill the s and p orbitals in the outer energy level. But, as you will see, if the electro-negativities are high enough and both atoms unwilling to give up electrons, sometimes atoms can deviate from and not follow) the Octet Rule.

Prelab:

Define following terms in your lab note book:

- VSEPR
- 2. Electronegativity
- 3. Octet Rule
- 4. Exception to octet rule
- 5. Electron domain

- 6. Electron geometry
- 7. Molecular geometry
- 8. Bond dipole
- 9. Molecule polarity

Molecule Shapes

Log on to https://phet.colorado.edu/en/simulation/molecule-shapes either by Googling "phet simulations molecule shape." Click on "Run Now." If a screen pops up asking to update Java, click on "Later."

Part 1: Electron domains

Explore the Model screen of the simulation and check what happens to existing atoms/lone pairs

- a. when you add an atom or lone pair
- b. when you remove an atom or lone pair.
- c. does adding an atom have same effect as adding a lone pair?

Part 2: 3D Drawings

- Line: In the plane of the paper: ——
- Wedge: Coming forward, in front of the plane of the paper: -
- Dash: Going backward, behind the plane of the paper: """
- Draw the molecule of CHFCIBr on left in your notebook and label the position of each atom as
 - In the plane of the paper
 - In front of the plane of the paper ______
 - Behind the plane of the paper ______



 Using the Model screen, add bonding domains (*) to the central atom (0). Using lines, wedges and dashes, draw each molecule's shape. Draw table in your lab notebook. Note: Bonding domains include any kind of bond and lone pairs.

Bonding Domains Around Central Atom	Drawing of Shape	Electron Geometry	Bond Angeles
2	e-o-e	Linear	180*
3	0		

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