Calculating pH of Acids - Practice

Instructions: Use the equations for pH to complete the problems below.

1) What is the pH of a 0.0266 M hydrobromic acid (HBr) solution?

$$pH = -log[H^+] \implies pH = -log[0.0266] \implies pH = 1.58$$

2) Calculate the hydrogen ion concentration of a solution with a pH of 6.3.

$$[H^+] = 10^{-pH} \longrightarrow [H^+] = 10^{-6.3} \longrightarrow [H^+] = 5.01 \times 10^{-7}$$

3) Calculate the pH of a 0.004 M hydrofluoric acid (HF) solution?

$$pH = -log[H^+] \implies pH = -log[0.004] \implies pH = 2.40$$

4) What is the hydrogen ion concentration of an acid with a pH of 2.21?

$$[H^+] = 10^{-pH} \longrightarrow [H^+] = 10^{-2.21} \longrightarrow [H^+] = 6.17 \times 10^{-3}$$

Calculate the pH of a 1.14 L solution that contains 0.19 moles of HCl.

$$[H^+] = \frac{\text{moles of solute}}{\text{liters of solvent}} \longrightarrow \frac{0.19 \text{ mol}}{1.14 \text{ L}} \longrightarrow 0.167 \text{ M}$$

$$pH = -log[H^+] \implies pH = -log[0.167] \implies pH = 0.78$$

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Arshad Iqbal

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