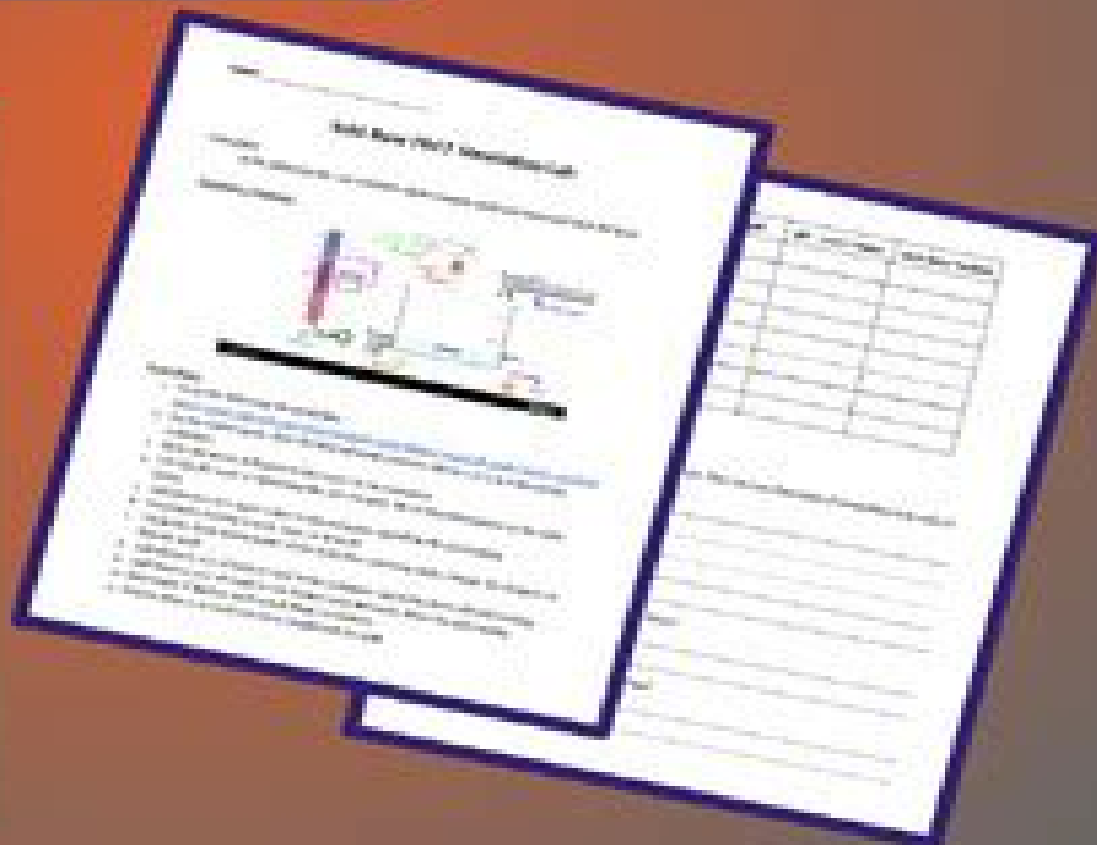


ACID-BASE PHET SIMULATION LAB

Included:

- Editable Worksheet
- Teacher's Key
- Teacher Suggestions



Phet Lab Acid And Base Packet

Robert De Levie



Phet Lab Acid And Base Packet:

American Druggist Blue Price Book, 1961 **Acids and Bases** Brian G. Cox, 2013-01-31 Acids and bases are ubiquitous in chemistry. Our understanding of them however is dominated by their behaviour in water. Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates. This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents. The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented. Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution. Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial. Experimental methods for determining dissociation constants are described in detail. In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed: protic solvents such as alcohols which are strong hydrogen bond donors; basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran. Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change. Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described. **Aqueous Acid-base Equilibria and Titrations** Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry. It is in addition a fresh and modern approach to a topic all too often taught in an outmoded way. This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace. Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered. The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria. Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH. In the description of acid base titrations a general master equation is developed. It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data. Graphical estimates of the steepness of titration curves are also

developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved

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