

# **ACTIVITY: Molecular Polarity Computer Simulation (PhET)**

Go to: <https://phet.colorado.edu/> Click on the arrow, DOWNLOAD, OPEN and RUN the MOLECULE POLARITY program. Or google "phet molecule polarity" to find the simulation.

## **Introduction:**

In this atomic level simulation, you will investigate how atoms' **electronegativity** values affect the bonds they produce. When two atoms form a covalent bond, a pair of electrons is shared between atoms. **Electronegativity** is a measure of a single atom's ability to attract electrons shared in that bond. In this lab you will work at your own pace, to answer a number of questions. To begin, from what you've already learned about the protons and electrons in an atom, what would cause an atom to have a **high** electronegativity value?

Why might an atom have a low electronegativity value?

## **Procedure:**

**New Molecule**

- ☛ Turn on (check) all view options. **distal** can give the molecule with your mouse.
- ☛ Take your time and investigate how the binary compound's bond behaves when the atom's electronegativity and orientation are changed. **Do not rush through this step.** Describe the bond formed between two atoms with **similar, low** electronegativities.

Describe the bond formed between two atoms with **similar, high** electronegativities.

Describe the bond formed between two atoms with **very different** electronegativities.

Describe (in your own words) what is meant by **partial charges**,  $\delta^-$  and  $\delta^+$ .

$\delta^-$  represents: \_\_\_\_\_

$\delta^+$  represents: \_\_\_\_\_

What happens when the electric field is applied to a very polar molecule?

Why do you think this is? \_\_\_\_\_

**View**

- ☒ Bond Dipoles **ON**
- ☐ Partial Charges
- ☐ Bond Orientation

**Surfaces**

- ☐ None
- ☐ Electrostatic Potential
- ☐ Electron Density

**Electric Field**

- ☐ On ☐ Off



# Phet Molecular Polarity Lab Answer Key

**Nivaldo J. Tro**



## **Phet Molecular Polarity Lab Answer Key:**

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## **Table of Contents Phet Molecular Polarity Lab Answer Key**

1. Understanding the eBook Phet Molecular Polarity Lab Answer Key
  - The Rise of Digital Reading Phet Molecular Polarity Lab Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Phet Molecular Polarity Lab Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Phet Molecular Polarity Lab Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Phet Molecular Polarity Lab Answer Key
  - Personalized Recommendations
  - Phet Molecular Polarity Lab Answer Key User Reviews and Ratings
  - Phet Molecular Polarity Lab Answer Key and Bestseller Lists

5. Accessing Phet Molecular Polarity Lab Answer Key Free and Paid eBooks
  - Phet Molecular Polarity Lab Answer Key Public Domain eBooks
  - Phet Molecular Polarity Lab Answer Key eBook Subscription Services
  - Phet Molecular Polarity Lab Answer Key Budget-Friendly Options
6. Navigating Phet Molecular Polarity Lab Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Phet Molecular Polarity Lab Answer Key Compatibility with Devices
  - Phet Molecular Polarity Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Phet Molecular Polarity Lab Answer Key
  - Highlighting and Note-Taking Phet Molecular Polarity Lab Answer Key
  - Interactive Elements Phet Molecular Polarity Lab Answer Key
8. Staying Engaged with Phet Molecular Polarity Lab Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Phet Molecular Polarity Lab Answer Key
9. Balancing eBooks and Physical Books Phet Molecular Polarity Lab Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Phet Molecular Polarity Lab Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Phet Molecular Polarity Lab Answer Key
  - Setting Reading Goals Phet Molecular Polarity Lab Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Phet Molecular Polarity Lab Answer Key
  - Fact-Checking eBook Content of Phet Molecular Polarity Lab Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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