



Trace Elliot Ah1000 Manual

JA Banks



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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b ba2 c br 2 d ca e ga3 f h 2 g h

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a i2o5
5co i2 5co2 i 2 o 5 5 c o i 2 5 c o 2

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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